

### Chapter 1: Stoichiometry

Q4.

Sol:  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$

Since 1 mole of  $\text{CaCO}_3$  produces 1 mole of  $\text{CaO}$ , So

1 mole of  $\text{CaO}$  will produce 1 mole of  $\text{CaCO}_3$ .

$\therefore$  1 mole of  $\text{CaO}$   $\rightarrow$  1 mole of  $\text{CaCO}_3$ .

Sol:  $\text{BaCl}_2 \rightarrow \text{BaSO}_4$

Given:

5 moles of  $\text{BaCl}_2$  contains 5 moles of  $\text{Ba}$  & 10 moles of  $\text{Cl}$  & each form 1 mole of  $\text{BaSO}_4$ ,  $\therefore$  each 1 mole of  $\text{BaCl}_2$  contains 1 mole of  $\text{Ba}$ .

Sol:  $2 \text{Na} \rightarrow \text{Na}_2\text{O}$

5 moles  $\text{O}_2$  given:

$\therefore$  2 moles produces 1 mole of  $\text{Na}_2\text{O}$

$\therefore$  5  $\rightarrow$   $\frac{5}{2}$

$\therefore$  5  $\rightarrow$   $\frac{5}{2} \times 2 = 5$  moles of  $\text{Na}_2\text{O}$ .

Q1.

Ans:  $\text{Ca} + \frac{1}{2} \text{O}_2 \rightarrow \text{CaO}$  (balanced chemical equation)

Wt. of  $\text{Ca} = 40 \text{ g/mole}$

So, moles of  $\text{Ca} = \frac{80}{40} = 2 \text{ moles}$

From above balanced chemical equation: 1 moles of  $\text{Ca}$  reacts with  $\frac{1}{2}$  moles of  $\text{O}_2$ .

Wt. of  $\text{O}_2$  with which 80g of  $\text{Ca}$  react  $= \frac{1}{2} \times 32 = 16 \text{ g}$

Ans.

Q2.

Sol:  $\text{NaCl} \rightarrow \text{NaCl}$

50 gm

If we have not balanced the chemical equation, then we have to apply P.M.C. rule or g-equivalent rule.

By P.M.C. rule (I) above

Mass of  $\text{Cl}$  absorbed by the reaction = Mass of  $\text{Cl}$  absorbed by the reaction

Mass of moles of  $\text{NaCl} = \text{Wt.} = 50 \text{ g}$  & mass of moles of  $\text{Cl}_2 = 71 \text{ g}$

$\frac{50}{58.5} = 2 \text{ moles of } \text{NaCl}$

moles of  $\text{Cl}_2 = \frac{50}{71}$

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**Scott C. Dulebohn**



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