At a high temperature, 0.50 mol of HBr was placed in a 1.0 L container and allowed to decompose according to the reaction:

$$2HBr(g) \rightleftharpoons H_{2(g)} + Br_{2(g)}$$

At equilibrium the [Br₂] was measured to be 0.13 M. What is K_{eq} for this reaction at this temperature?

	2HBr =	= H2 +	Br2
[x]	0.50	0	0
[C]	-0-26	+0.13	+0.13
IE]	0.24	0.13	0.13

Keg =	IH2][Br]2	= (0.13)2	=0.29
	100001100000000000000000000000000000000	150000000000000	

Answer Keg = 0.29

 When 1.0 mol of NH_{3(g)} and 0.40 mol of N_{2(g)} are placed in a 5.0 L vessel and allowed to reach equilibrium at a certain temperature, it is found that 0.78 mol of NH₃ is present. The reaction is:

2NH3(g) = 3H2(g) + N2(g)

Initial [NH3] = 1.0 mol = 0.20 M Initial [N2] = 0.40 mol = 0.080 M

Equil [NH3] = 0.78 mol = 0.156 M 2NH2 = 3 H2 + N2

	aNH3 &	= 3 Na	+ N2
[I]	0.20	0	0.080
EC3	-0.044	+0.066	+0.022
LEJ.	0.156	0.066	0.102

a) Calculate the equilibrium concentrations of all three species.

[NH₃] - 0.16 M [H₂] - 0.066 M [N₂] - 0.10 M

b), Calculate the value of the equilibrium constant at this temperature. (use unrounded cones, then round to 2503)

 $\text{Keq} = \frac{\text{EH}_2 \text{J}^3 \text{EN}_2 \text{J}}{\text{ENH}_3 \text{J}^2} = \frac{(0.066)^3 (0.102)}{(0.156)^2} = 0.0012$ Answ

Answer 1.2×10-3

c) How many moles of H₂ are present at equilibrium?

0.066 M x 5.0 L = 0.33 mal

Answer 0.33 mol

d) How many moles of N₂ are present at equilibrium?

0.102 M x 5.0 L = 0.51 mal

Answer 0.51 mol

Modeling Chemistry Ws Answers

Karin Nielsen-Saines

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- 1. Understanding the eBook Modeling Chemistry Ws Answers
 - The Rise of Digital Reading Modeling Chemistry Ws Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Modeling Chemistry Ws Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Ws Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modeling Chemistry Ws Answers
 - Personalized Recommendations
 - Modeling Chemistry Ws Answers User Reviews and Ratings
 - Modeling Chemistry Ws Answers and Bestseller Lists
- 5. Accessing Modeling Chemistry Ws Answers Free and Paid eBooks
 - Modeling Chemistry Ws Answers Public Domain eBooks
 - Modeling Chemistry Ws Answers eBook Subscription Services
 - o Modeling Chemistry Ws Answers Budget-Friendly Options
- 6. Navigating Modeling Chemistry Ws Answers eBook Formats

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- Modeling Chemistry Ws Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Modeling Chemistry Ws Answers
 - Interactive Elements Modeling Chemistry Ws Answers
- 8. Staying Engaged with Modeling Chemistry Ws Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Ws Answers
- 9. Balancing eBooks and Physical Books Modeling Chemistry Ws Answers
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Modeling Chemistry Ws Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Ws Answers
 - Setting Reading Goals Modeling Chemistry Ws Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Ws Answers
 - Fact-Checking eBook Content of Modeling Chemistry Ws Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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