

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out the products of the reaction. Show us the products and how to find them. The next 4 problems cover three types.

1. Find the mass of product of reaction of potassium hydroxide, KOH , and HCl , using stoichiometry. (a) potassium hydroxide, KOH , and HCl react to form potassium chloride.

Equation:	$\text{KOH} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2\text{O}$
Balance:	1:1:1:1
Amount:	4.00 10.0 1.00
Mass:	56 36.5 74.5

1.00 moles of
 KOH gives 1.00

2. How much a displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How much of hydrochloric acid is required to produce 4.00 g of magnesium?

Equation:	$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
Balance:	1:2:1:1
Amount:	10.00 20.00 10.00 2.00
Mass:	24 73 95 2

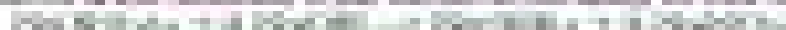
10.00 moles of
 Mg gives 20.00

3. Hydrogen, H_2 , reacts with oxygen, O_2 , to produce water. How much gas will react to give 10.00 g of hydrogen gas by weighing as much as the other gas? (How much of it is there?)

Equation:	$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
Balance:	2:1:2
Amount:	10.00 5.00 10.00 18.00
Mass:	4 32 18

10.00 moles of H_2 gives 5.00 moles of O_2 , or 160.00 g

4. Chlorine gas reacts with sodium hydroxide to form sodium chloride and a chlorine compound. How much sodium hydroxide is required to give 10.00 g of chlorine compound?



10.00 g of Cl_2O is 0.100 moles

Equation:	$\text{Cl}_2 + 2\text{NaOH} \rightarrow 2\text{NaCl} + \text{Cl}_2\text{O}$
Balance:	1:2:2:1
Amount:	10.00 20.00 20.00 10.00
Mass:	71 80 117 85

10.00 moles of Cl_2O is 0.100 moles of Cl_2

Modeling Chemistry Unit 8 Review

Christian G. Meyer



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