

## Unit 2 Worksheet 1: Molar relationships

### For each of the problems below:

- Write a full balanced chemical equation.
- Calculate, where it applies, molar masses and select the most useful molar ratio.
- How much of a mass balanced experimental demonstration exists in different units for each of the

1. Hydrogen and nitrogen, which react in the reaction  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ , are the products of the reaction and react. How many moles of nitrogen gas would be needed to completely react 1 mole of hydrogen gas?

Equation:	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
<b>Reactants</b>	1	3	1	3
<b>Products</b>	2	2	2	2

**Answer:**  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$  or  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$

2. Propane,  $\text{C}_3\text{H}_8$ , reacts in air to form carbon dioxide and water. Write a balanced reaction equation and calculate how many moles of propane are burned?

Equation:	$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$	$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$	$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$	$\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
<b>Reactants</b>	1	5	1	5
<b>Products</b>	3	4	3	4

**Answer:**  $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$  or  $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

3. Ammonia,  $\text{NH}_3$ , the nitrogen is made by reacting hydrogen and nitrogen in a reaction with nitrogen gas and hydrogen. How many moles of a substance are the reactants from the reaction of hydrogen gas?

Equation:	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
<b>Reactants</b>	1	3	1	3
<b>Products</b>	2	2	2	2

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**Ian Pickup**



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