

For the reverse reaction?

decreased the $E_a(r)$

- q) What effect did the catalyst have on the ΔH of the forward reaction? no effect

The reverse reaction?

no effect

12. Name four instances in which *catalysts* are used in industry or everyday life and tell which *catalysts* are used.

V_2O_5 used in making sulphuric acid (step $2SO_2 + O_2 \xrightarrow{V_2O_5} 2SO_3$)

maltase breaks down maltose to simple sugar (glucose)

nickel used to hydrogenate unsaturated fats

Pt, Pd, Rh used in catalytic converters to break down NO_2 in exhaust

13. Describe what happens to the *kinetic energy*, *potential energy* and the *total energy* of reactant molecules as they approach each other.

KE decreases PE increases TE remains constant.

14. Explain why a lower *activation energy* for a reaction leads to a greater reaction rate at a given temperature.

molecular collisions don't require as high an energy to be successful so at a particular temp. a greater fraction of the molecules will have enough energy.

15. A small piece of zinc react with 2.0 M HCl to produce 12.0 mL of H_2 gas in 30.0 seconds. Calculate the *rate of reaction*

a) In mL of H_2 /second $\frac{12.0 \text{ mL } H_2}{30.0 \text{ s}} = 0.400 \text{ mL } H_2/\text{s}$

Answer $0.400 \text{ mL } H_2/\text{s}$

b) In moles of H_2 /second

$$\frac{0.400 \text{ mL } H_2}{\text{s}} \times \frac{1 \text{ L}}{1000 \text{ mL}} \times \frac{1 \text{ mol}}{22.4 \text{ L}} = 1.79 \times 10^{-5} \text{ mol } H_2/\text{s}$$

Answer $1.79 \times 10^{-5} \text{ mol } H_2/\text{s}$

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Annelies Wilder-Smith



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