

Stoichiometry Worksheet 2: Percent Yield

What aspects of the procedure do you think:

- affected the theoretical or actual yield negatively?
- substantially affected your yield in a positive or negative way? explain how these things can be made up.
- affected your setup or the analysis. (Cleaning up flasks, glass break, incomplete transfer, incorrect mass balance)

10. Using the theoretical mass of the silver chloride, calculate the theoretical mass of silver chloride. Show the calculation. What mass of silver chloride was actually obtained? (Percent yield)

$$\text{Mass AgNO}_3 = 1.0 \text{ g} \times \frac{143.32 \text{ g AgCl}}{169.87 \text{ g AgNO}_3} = 0.843 \text{ g AgCl}$$

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$$\text{Mass AgCl} = \frac{\text{Mass AgNO}_3 \times \text{Molar Mass AgCl}}{\text{Molar Mass AgNO}_3}$$

$$\text{Mass AgCl} = \frac{1.0 \text{ g} \times 143.32 \text{ g/mol}}{169.87 \text{ g/mol}} = 0.843 \text{ g}$$

$$\text{Mass AgCl} = \frac{\text{Mass AgCl}}{\text{Mass AgCl}} = 1.0$$

11. Calculate the percent yield of silver chloride. Show the calculation. What mass of silver chloride was actually obtained? (Percent yield)

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12. What mass of silver chloride was actually obtained? (Percent yield) Calculate the percent yield of silver chloride. Show the calculation. What mass of silver chloride was actually obtained? (Percent yield)

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13. Calculate the percent yield of silver chloride. Show the calculation. What mass of silver chloride was actually obtained? (Percent yield)

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Modeling Chemistry U8 Ws 1 V2 0

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Modeling Chemistry U8 Ws 1 V2 0:

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