For the reverse reaction? <u>(lec/lased</u> the <u>Ea/r)</u>
q) What effect did the catalyst have on the ΔH of the forward reaction? ho effect
The reverse reaction? no effect
12. Name four instances in which catalysts are used in industry or everyday life and tell which catalysts are used. V20s Used in making sulphuric acid/step 2502 + 02 2503
TOS THE PROPERTY OF THE PROPER
mattase breaks down mattose to simple sugar (glucose)
nickel used to hydrogenate unsaturated fats
Pt, Pd, Rh used in catalytic converters to break down NO2 in exhaus
 Describe what happens to the kinetic energy, potential energy and the total energy of reactant molecules as they approach each other.
KE decreases PE increases TE remains constant.
 Explain why a lower activation energy for a reaction leads to a greater reaction rate at a given temperature.
molecular collisions of require as high an enormy to be surrounded
molecular collisions of require as high an energy to be successful so of a particular temp. a greater fraction of the molecules will have enough energy. A small piece of this react with 2.0 M HCI to produce 12.0 mL of Hz was in 30.0 seconds.
Enduck Energy 15. A small piece of zitic react with 2.0 M HCl to produce 12.0 mL of H ₂ gas in 30.0 seconds. Calculate the rate of reaction
a) In ml. of Ho/second 12.0 mL H2 = 0.400 mL H= /s
a) In mL of H ₂ /second 12.0 mL H ₂ = 0.400 mL H ₂ /s 30.0 s
30.05 Answer 0.400 mL Hg/s
b) In moles of Hy/second
0.400 m L Hz × 1 L : 1 mol = 1.79 × 10 5 mol Hz/s Answer 1.79 × 10 5 mol Hz/s
1.79×10 5 mol H2/s

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S Ashworth

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