

Chemistry – Unit 5 Worksheet 1

Relative Abundance of the Elements

Relative Mass From Counts

We can estimate the relative abundance of elements by counting the number of atoms of each element present in a sample.

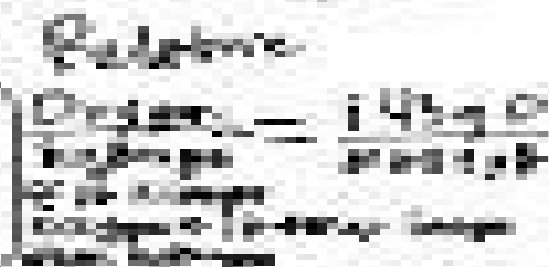
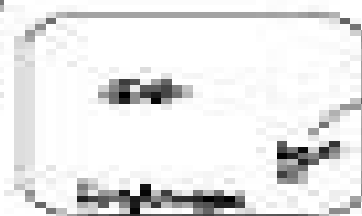
1. Equal volumes of gases contain the same number of molecules as the other gases present.

Example: **Hydrogen**

2. Atoms of the same element are going to have the same mass.

H₂, O₂, N₂, Ar, CO₂

1 cm³ of each gas at the same pressure and temperature has the same number of molecules. But the density of oxygen gas is 1.43 g/cm³ and hydrogen gas is 0.09 g/cm³. This shows that oxygen is 16 times heavier than hydrogen. The relative abundance of hydrogen is **1/16**.



Chemists have found that elements combine in the same way as the relative mass of all the elements using this technique. Measurements of the density of the gases present in many other elements are difficult, it was impossible. However, it is now possible to determine relative abundance using mass spectrometry, which is a different method.

Relative Mass From Compounds

Many compounds contain equal numbers of atoms of each element. For example, water (H₂O) contains two atoms of hydrogen and one atom of oxygen. The relative mass of water is 18 g/mol. The relative mass of hydrogen is 2 g/mol. The relative mass of oxygen is 16 g/mol.

For example, if you have 18 g of water, you have 2 g of hydrogen and 16 g of oxygen. This is because the relative mass of water is 18 g/mol. The relative mass of hydrogen is 2 g/mol. The relative mass of oxygen is 16 g/mol.



Modeling Chemistry U5 Ws1 V3 0

Mark Steyvers



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