

Chemistry – Unit 5 Worksheet 1

Relative Abundance of the Elements

Relative Mass From Counts

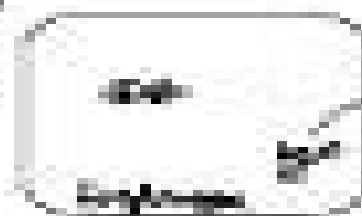
We have established that the relative abundance of an element is the mass of one representative atom.

1. Equal numbers of atoms contain the same number of nucleons as the same physical mass of elements. **Simplest Equation**
2. Number of nucleons of one atom is given as the relative abundance.

Ex. ${}^2\text{H}$, ${}^3\text{H}$, ${}^4\text{He}$, ${}^6\text{Li}$, ${}^{12}\text{C}$, ${}^{16}\text{O}$

Using the periodic table, we can determine the relative abundance of each element. The relative abundance of an element is the mass of one representative atom. The relative abundance of an element is the mass of one representative atom. The relative abundance of an element is the mass of one representative atom.

Oxygen $2 \times 16 = 32$ 1 Liter	Hydrogen $2 \times 1 = 2$ 1 Liter
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Relative Oxygen $2 \times 16 = 32$ 1 Liter	Hydrogen $2 \times 1 = 2$ 1 Liter
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It is important to remember that the relative abundance of an element is the mass of one representative atom. The relative abundance of an element is the mass of one representative atom. The relative abundance of an element is the mass of one representative atom.

Relative Mass From Composition

Many elements combine to form compounds. The relative abundance of an element is the mass of one representative atom. The relative abundance of an element is the mass of one representative atom. The relative abundance of an element is the mass of one representative atom.

For example, the relative abundance of an element is the mass of one representative atom. The relative abundance of an element is the mass of one representative atom. The relative abundance of an element is the mass of one representative atom.



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