

Chemistry Unit 7 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedures outlined in class. Be sure to circle your final answer.

- Calculate the number of moles of potassium chlorate, KClO_3 , that must decompose to produce potassium chloride, KCl , and 1.0 moles of oxygen gas.
- In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 0.40 g of magnesium?
- Ethanol, $\text{C}_2\text{H}_5\text{OH}$, reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.00 moles of ethanol?
- Determine the mass of sodium chloride produced when 0.70 g of sodium (Na) reacts with sodium hydroxide according to the following equation:
 $\text{NaClO}_2 + 2 \text{NaOH} \rightarrow \text{NaClO}_3 + 2 \text{NaOH}$
- In the copper-silver nitrate lab copper metal and silver nitrate solution reacted to produce silver metal and copper(II) nitrate in solution.
 A student placed a copper wire with a mass of 0.05 g in the reaction test tube. The silver-silver nitrate solution contained 0.40 g of silver nitrate.
 She obtained 0.07 g of silver metal. Calculate the percent yield of silver.
- When hydrochloric acid (HCl) is added to sodium hydroxide solution, the products are water, aqueous sodium chloride and sodium chloride gas. What is the per cent yield of 0.04 g of Cl_2 gas collected when 0.01 g of sodium hydroxide solution reacts with excess HCl?
- Phosphorus and bromine react (phosphoryl bromide) to form phosphorus tribromide. If 0.01 g of phosphorus and 0.01 g of bromine react, how many grams of PBr_3 could be produced?
- Iron sulfate and oxygen gas react to form iron oxide and sulfur dioxide. Determine the amount of SO_2} that should be produced in a reaction between 0.01 g of FeSO_4 and 0.01 g of oxygen. What is the percent of the experiment?

$$1. \quad 1.2 \text{ moles of } \text{KClO}_3 \quad 2. \quad 0.0001 \text{ moles of } \text{H}_2\text{O} \quad 3. \quad 0.0001 \text{ g } \text{NaCl} \quad 4. \quad 0.0001 \text{ g of } \text{NaClO}_3$$

$$5. \quad 0.01\% \quad 6. \quad 0.0001\% \quad 7. \quad 0.01\% \text{ of } \text{PBr}_3 \quad 8. \quad 0.0001 \text{ g of } \text{SO}_2 \quad 9. \quad 0.0001 \text{ g of } \text{SO}_2$$

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