

## CP Chemistry Unit 1 Test Review KEY

### Problems and questions:

- Round each of the following numbers to 3 significant figures:  
a) .09882 **.0988**      b) 123,459 **123,000**      c) 102.45 **102**      d)  $2.301 \times 10^{-2}$   **$2.30 \times 10^{-2}$**
- Calculate, use sig figs:  $(12.22 \times 14.02 \times 120.1 \times .555) / 10.00 =$  **1140**
- A sample of metal having a mass of 12.122 g is added to a graduated cylinder containing 35.2 mL of water. The water level rises to 38.4 mL. Calculate the density, use correct significant figures.  
$$d = \frac{12.122}{3.2} = 3.8 \text{ g/mL}$$
- Calculate the % error for an experiment where the student measured 3.45 g and the actual value was 3.00 g?  
$$\% \text{ error} = \frac{(3.45 \text{ g} - 3.00 \text{ g})}{(3.00 \text{ g})} \times 100 = 15.0\%$$
- Write the symbol for the following ions: Mg losing 2 e<sup>-</sup>, Al losing 3 e<sup>-</sup>, N gaining 3 e<sup>-</sup>, S gaining 2 e<sup>-</sup>  
**Mg<sup>2+</sup>, Al<sup>3+</sup>, N<sup>3-</sup>, S<sup>2-</sup>**
- a) 2 b) 3 c) 12 d) 5
- List three points of Dalton's atomic theory.  
A. Elements made of tiny particles, called atoms.  
B. Atoms of a given element are identical; the atoms of different elements are different.  
C. In chemical reactions, atoms are combined, separated, or rearranged. Atoms cannot be subdivided, created, or destroyed.  
D. Atoms combine with atoms, in small whole numbers, to form compounds.
- Describe Thompson's experiment with the cathode ray tube, what did he discover?  
Cathode ray tube consists of a cathode (negative) and anode (positive) side. The "ray" could be bent with magnetic or electric fields. The "ray" bent toward a positive EM field and away from a negative EM field. Discovered electron and charge to mass ratio of electron.
- Describe Millikan's Oil Drop experiment. What exactly did he determine?  
He suspended a micro oil droplet in an electric field and was able to determine the actual charge of an electron to be  $-1.6 \times 10^{-19} \text{ C}$ . Also, determined the mass of an electron.
- Explain Rutherford's Gold Foil experiment. What were three observations and three conclusions?  
He fired alpha particles at gold foil and observed: 1) most went straight through, 2) small number showed a slight deflection, 3) very few bounced straight back. He concluded: 1) mass of atom is in the nucleus, 2) nucleus is positively charged, 3) atom mostly empty space.
- What is an isotope? **Atoms of the same element that differ by # of electrons.**
- Write the symbol of the most abundant copper isotope if the average atomic mass is 63.55?  **$^{64}_{29}\text{Cu}$**
- Write the nuclear symbol for an isotope with Z=40, and 45 neutrons?  **$^{85}_{40}\text{Zr}$**
- Consider the diagram and match the letters with the appropriate particle  
**A= electron, B= neutron, C= proton**
- Place the forms of hydrogen (H, H<sup>+</sup>, H<sup>-</sup>, H<sub>2</sub>) in the table.



Atom	Cation	Anion	Molecule
<b>H</b>	<b>H<sup>+</sup></b>	<b>H<sup>-</sup></b>	<b>H<sub>2</sub></b>

# Modeling Chemistry Unit 1 Review Answers Key

**Liying Dong**



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