

Hess's Law Practice Problems

Please check back for the answers. I will repost this document with the answers at a later date

1. Given the following equations and ΔH° values, determine the heat of reaction (kJ) at 298 K for the reaction:



$$\Delta H^\circ/\text{kJ} = -112.5$$



$$\Delta H^\circ/\text{kJ} = -493.4$$



$$\Delta H^\circ/\text{kJ} = -92.3$$

Answer: -1376 kJ

2. Given the following equations and ΔH° values, determine the heat of reaction (kJ) at 298 K for the reaction:



$$\Delta H^\circ/\text{kJ} = -276.6$$



$$\Delta H^\circ/\text{kJ} = -827.5$$



$$\Delta H^\circ/\text{kJ} = -296.9$$

Answer: -319.5 kJ

3. Determine $\Delta H^\circ/\text{kJ}$ for the following reaction using the listed enthalpies of reaction:

Hess Law Practice Answers

Valentine Korah



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