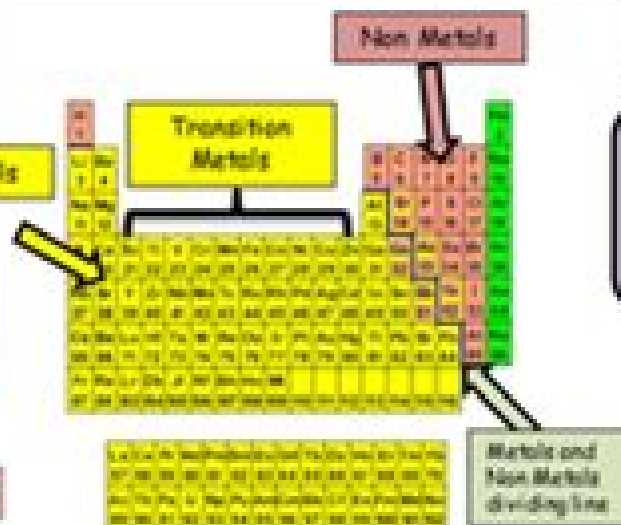


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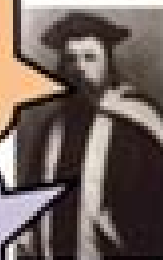


Particle	Relative Mass	Relative Charge
Proton	1	+
Neutron	1	0
Electron	Very small	-



I came up with the first periodic table. I arranged the elements in order of increasing atomic mass

Dmitri Mendeleev



I put elements with similar properties into horizontal rows. I also left gaps incase any new elements were discovered

He

4 ← **MASS NUMBER** = number of protons + number of neutrons

2 ← **PROTON NUMBER** = number of protons (atomic)

CHARGE

Isotopes

Higher

Example is an atom with a different number of neutrons

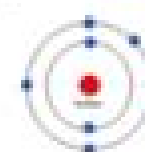
Notice that the mass number is different. How many neutrons does each isotope have?

16 17 18

Each isotope has 8 protons - if it didn't then it just wouldn't be oxygen any more

N

14
7



Electron structure
+ 2, 5

Mg

24
12



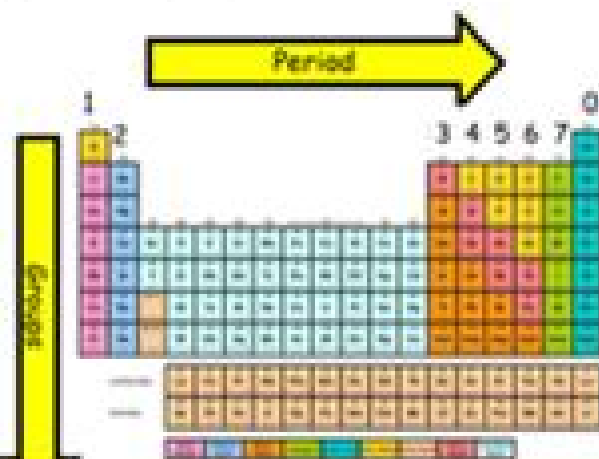
Electron structure
+ 2, 8, 2

Ca

40
20



Electron structure
+ 2, 8, 8, 2



How can an atom have a **decimal** for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 so the average atomic mass is:

$$35 \times 3 + 1 \times 37 = 106$$

Elements in the same group have the same number of electrons in the outer shell (this corresponds to their group number)



(Most important) All of the elements in the same group have similar **PROPERTIES**. This is how I thought of the periodic table in the first place. This is called **PERIODICITY**.

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Rod Beavon



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1. Understanding the eBook Gcse Edexcel Chemistry C2 2013
 - The Rise of Digital Reading Gcse Edexcel Chemistry C2 2013
 - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Edexcel Chemistry C2 2013
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gcse Edexcel Chemistry C2 2013
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Edexcel Chemistry C2 2013
 - Personalized Recommendations
 - Gcse Edexcel Chemistry C2 2013 User Reviews and Ratings
 - Gcse Edexcel Chemistry C2 2013 and Bestseller Lists
5. Accessing Gcse Edexcel Chemistry C2 2013 Free and Paid eBooks
 - Gcse Edexcel Chemistry C2 2013 Public Domain eBooks
 - Gcse Edexcel Chemistry C2 2013 eBook Subscription Services

- Gcse Edexcel Chemistry C2 2013 Budget-Friendly Options
- 6. Navigating Gcse Edexcel Chemistry C2 2013 eBook Formats
 - ePub, PDF, MOBI, and More
 - Gcse Edexcel Chemistry C2 2013 Compatibility with Devices
 - Gcse Edexcel Chemistry C2 2013 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gcse Edexcel Chemistry C2 2013
 - Highlighting and Note-Taking Gcse Edexcel Chemistry C2 2013
 - Interactive Elements Gcse Edexcel Chemistry C2 2013
- 8. Staying Engaged with Gcse Edexcel Chemistry C2 2013
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gcse Edexcel Chemistry C2 2013
- 9. Balancing eBooks and Physical Books Gcse Edexcel Chemistry C2 2013
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gcse Edexcel Chemistry C2 2013
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gcse Edexcel Chemistry C2 2013
 - Setting Reading Goals Gcse Edexcel Chemistry C2 2013
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gcse Edexcel Chemistry C2 2013
 - Fact-Checking eBook Content of Gcse Edexcel Chemistry C2 2013
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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