The equation for another reaction used in industry is

• 
$$CO(g) + H_2O(g) \rightleftharpoons H_2(g) + CO_2(g)$$
  $\Delta H^0 = -42 \text{ kJ}$ 

- Under certain conditions of temperature and pressure, 2.0 mol of carbon monoxide and 3.2 mol of steam were left to reach equilibrium. At equilibrium, 1.6 mol of both hydrogen and carbon dioxide were present.
  - Calculate the amounts of carbon monoxide and steam at equilibrium and the value of K<sub>c</sub>

# **Equilibrium Problems Ice Method**

**SA Adler** 

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