

Empirical and Molecular Formula Worksheet

SHOW WORK ON A SEPARATE SHEET OF PAPER.

Write the empirical formula for the following compounds.

- 1) C_6H_6
- 2) C_8H_{18}
- 3) WO_2
- 4) $C_2H_6O_2$
- 6) A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole. What is the molecular formula of this compound?
- 7) A compound with an empirical formula of C_4H_4O and a molar mass of 136 grams per mole. What is the molecular formula of this compound?
- 8) A compound with an empirical formula of $CFBrO$ and a molar mass of 254.7 grams per mole. What is the molecular formula of this compound?
- 9) A compound with an empirical formula of C_2H_8N and a molar mass of 46 grams per mole. What is the molecular formula of this compound?
- 11) Nitrogen and oxygen combine to form a compound that is 36.84% N. What is the empirical formula.
- 12) A sample of indium chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the indium compound?
- 13) An unknown compound was found to have a percent composition as follows: 47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula? If the true molar mass of the compound is 166.22 g/mol, what is its molecular formula?
- 14) Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rubbing alcohol?

Empirical And Molecular Formula Worksheet Answer Key

LL Leslie



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