

Chemistry – Unit 5 Worksheet 3

Empirical and Molecular Formulas

Show all your work when solving the following problems. Be sure to include units and label your answer.

1. Find the empirical formula of a compound containing 32.0 g of bromine and 4.9 g of magnesium.

$$32.0 \text{ g} \times \frac{1 \text{ mole}}{79.9 \text{ g}} = 0.401 \text{ mole Br}$$

1:2 ratio

$$4.9 \text{ g} \times \frac{1 \text{ mole}}{24.31 \text{ g}} = 0.20 \text{ mole Mg}$$

 MgBr_2

2. What is the empirical formula of a carbon-oxygen compound, given that a 95.2 g sample of the compound contains 40.8 g of carbon and the rest oxygen?

$$95.2 \text{ g} - 40.8 \text{ g} = 54.4 \text{ g O}$$

$$40.8 \text{ g} \times \frac{1 \text{ mole}}{12.0 \text{ g}} = 3.40 \text{ moles C}$$

1:1 ratio

$$54.4 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 3.40 \text{ moles O}$$

 CO

3. A compound was analyzed and found to contain 9.8 g of nitrogen, 0.70 g of hydrogen, and 33.6 g of oxygen. What is the empirical formula of the compound?

$$9.8 \text{ g} \times \frac{1 \text{ mole}}{14.0 \text{ g}} = 0.70 \text{ moles N}$$

1:1:3 ratio

$$0.70 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 0.69 \text{ moles H}$$

 HNO_3

$$33.6 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 2.10 \text{ moles O}$$

4. A compound composed of hydrogen and oxygen is found to contain 0.59 g of hydrogen and 9.40 g of oxygen. The molar mass of this compound is 34.0 g/mol. Find the empirical and molecular formulas.

$$0.59 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 0.58 \text{ mole H}$$

1:1 ratio; empirical formula: HO
 empirical formula mass: 17.0 g/mol
 molecular formula mass: 34.0 g/mol
 mfm is 2x efm so the molecular formula is: H_2O_2

$$9.40 \text{ g} \times \frac{1 \text{ mole}}{16.0 \text{ g}} = 0.588 \text{ mole O}$$

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DP Hallahan



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1. Understanding the eBook Empirical And Molecular Formula Answer Key
 - The Rise of Digital Reading Empirical And Molecular Formula Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical And Molecular Formula Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical And Molecular Formula Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical And Molecular Formula Answer Key
 - Personalized Recommendations
 - Empirical And Molecular Formula Answer Key User Reviews and Ratings
 - Empirical And Molecular Formula Answer Key and Bestseller Lists
5. Accessing Empirical And Molecular Formula Answer Key Free and Paid eBooks

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 - Empirical And Molecular Formula Answer Key eBook Subscription Services
 - Empirical And Molecular Formula Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Empirical And Molecular Formula Answer Key Compatibility with Devices
 - Empirical And Molecular Formula Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical And Molecular Formula Answer Key
 - Highlighting and Note-Taking Empirical And Molecular Formula Answer Key
 - Interactive Elements Empirical And Molecular Formula Answer Key
 8. Staying Engaged with Empirical And Molecular Formula Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical And Molecular Formula Answer Key
 9. Balancing eBooks and Physical Books Empirical And Molecular Formula Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical And Molecular Formula Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Empirical And Molecular Formula Answer Key
 - Setting Reading Goals Empirical And Molecular Formula Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Empirical And Molecular Formula Answer Key
 - Fact-Checking eBook Content of Empirical And Molecular Formula Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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


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