

Unit 6 – Thermodynamics (7-9%)

Changes in a substance's properties or change into a different substance requires an exchange of energy.

6.1 Endothermic and Exothermic Processes: Explain the relationship between experimental observations and energy changes associated with a chemical or physical transformation.

- Temperature changes in a system indicate energy changes.
- Energy changes in a system can be described as endothermic and exothermic processes such as the heating or cooling of a substance, phase changes, or chemical transformations.
- When a chemical reaction occurs, the energy of the system either decreases (exothermic reaction), increases (endothermic reaction), or remains the same. For exothermic reactions, the energy lost by the reacting species (system) is gained by the surroundings, as heat transfer from or work done by the system. Likewise, for endothermic reactions, the system gains energy from the surroundings by heat transfer to or work done on the system.
- The formation of a solution may be an exothermic or endothermic process, depending on the relative strengths of intermolecular/interparticle interactions before and after the dissolution process.

6.2 Energy Diagrams: Represent a chemical or physical transformation with an energy diagram.

- A chemical or physical process is represented with an energy diagram that shows the endothermic or exothermic nature of the process.

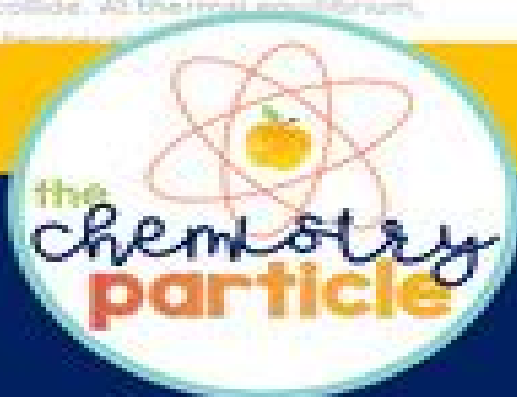
Unit 6 Objectives

Thermodynamics

- Particles in a warmer body have a greater average kinetic energy than those in a cooler body.
- Collisions between particles in a warmer body and a cooler body are called "thermal collisions." The process is called "thermalization."
- Eventually, thermal equilibrium is reached as the particles continue to collide. At thermal equilibrium, the average kinetic energy of both bodies is the same and there is no further thermalization.

AP Chemistry

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