

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that we are a "paper" writer, a chemist who has to write out what products may be made. We have the reactants and know the mass. The mass is underlined (underlined means given).

1. Chlorine gas reacts with sodium metal to produce sodium chloride. NaCl has a molar mass of 58.44 g/mol. How much sodium chloride is produced from 10.0 g of chlorine gas?

Equation:	$\text{Cl}_2 + 2\text{Na} \rightarrow 2\text{NaCl}$	1	2
Balance:	2 x 71	2 x 23	2
Amount:	10.0	10.0	10.0
Answer:	58.44	58.44	58.44

Answer:
58.44 g NaCl

2. How much a displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How much magnesium chloride is produced by reacting 10.0 g of magnesium?

Equation:	$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$	1	2
Balance:	24.31	2	2
Amount:	10.0	10.0	10.0
Answer:	58.44	58.44	58.44

Answer:
58.44 g MgCl_2

3. Sodium, Na , reacts with hydrochloric acid to produce sodium chloride gas and hydrogen. How much sodium chloride is produced by reacting 10.0 g of sodium?

Equation:	$2\text{Na} + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2$	2	2
Balance:	2 x 23	2 x 58.44	2
Amount:	10.0	10.0	10.0
Answer:	58.44	58.44	58.44

Answer: 58.44 g NaCl

4. Chlorine gas reacts with sodium metal to produce sodium chloride and hydrogen gas. How much sodium chloride is produced by reacting 10.0 g of chlorine gas?

Equation:	$\text{Cl}_2 + 2\text{Na} \rightarrow 2\text{NaCl}$	1	2
Balance:	2 x 71	2 x 23	2
Amount:	10.0	10.0	10.0
Answer:	58.44	58.44	58.44

Answer: 58.44 g NaCl

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G Orfield



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