

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that during "paper", writing a composition about how the world will change people may disagree. However, this problem is not based on this. The main 3 numbers (mass, molar mass, moles).

1. Calculate the number of moles of potassium hydroxide, KOH , and then, mass of potassium hydroxide, KOH , and 1.0 moles of potassium gas.

Equation:	$\text{KOH} \rightarrow \text{K} + \text{O}_2$	$\text{KOH} \rightarrow \text{K} + \text{O}_2$	$\text{KOH} \rightarrow \text{K} + \text{O}_2$
Substance:	1.0 g	1.0 g	1.0 g
Mass:	56.1	39.1	16.0
Moles:	0.018	0.026	0.063

1.0 moles of
KOH is 56.1 g

2. How much a displacement of potassium, potassium metal reacts with hydrochloric acid to produce potassium chloride and hydrogen gas. How much potassium chloride and hydrogen gas are produced by reacting 1.0 mole of potassium?

Equation:	$\text{K} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2$	$\text{K} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2$	$\text{K} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2$
Substance:	1.0 mole	1.0 mole	1.0 mole
Mass:	39.1	35.5	39.1
Moles:	1.0	1.0	1.0

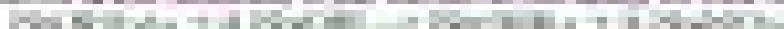
1.0 mole of
K is 39.1 g

3. Potassium, K , reacts with hydrochloric acid to produce potassium chloride and hydrogen gas. How much potassium chloride and hydrogen gas are produced by reacting 1.0 mole of potassium?

Equation:	$\text{K} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2$	$\text{K} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2$	$\text{K} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2$	$\text{K} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2$
Substance:	1.0 mole	1.0 mole	1.0 mole	1.0 mole
Mass:	39.1	35.5	39.1	39.1
Moles:	1.0	1.0	1.0	1.0

1.0 mole of K is 39.1 g

4. Potassium, K , reacts with hydrochloric acid to produce potassium chloride and hydrogen gas. How much potassium chloride and hydrogen gas are produced by reacting 1.0 mole of potassium?



1.0 mole of K is 39.1 g

Equation:	$2\text{K} + 2\text{HCl} \rightarrow 2\text{KCl} + \text{H}_2$	$2\text{K} + 2\text{HCl} \rightarrow 2\text{KCl} + \text{H}_2$	$2\text{K} + 2\text{HCl} \rightarrow 2\text{KCl} + \text{H}_2$
Substance:	1.0 mole	1.0 mole	1.0 mole
Mass:	39.1	35.5	39.1
Moles:	1.0	1.0	1.0

1.0 mole of K is 39.1 g

Modeling Chemistry U8 Test A V2 0

Yan Bai



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