

Unit 7 Chemistry - Nail Lab

Purpose: After you have completed the lab, think about what was the overall goal for performing the lab. What were we trying to accomplish?

Procedure:

Step 1

1. Mass a clean & labeled _____ mL beaker.
2. Put between _____ g of copper (II) chloride in the beaker. Record the mass of the beaker and the copper (II) chloride.
3. Add _____ mL of distilled water to the beaker. Stir to dissolve the solid.
4. Obtain the mass of _____ iron nails together to ~ 0.5 g.
5. Place the nails in the copper (II) chloride solution. Observe the reactions; record your observations. Place the labeled beaker in the place designated by your teacher.

Step 2

- a. Remove the nails. Rinse or scrape the _____ (copper metal) from the nails into your labeled beaker. Place the nails in a separate labeled small beaker. Note the appearance of the nails in your observations. Place this beaker in the drying oven.
- b. _____ the solution from the 250 mL beaker. Rinse the precipitate with about _____ mL of distilled water. Try to lose as little of the solid copper as you can when you decant. Add a 2nd rinse with distilled water, rinse the copper with _____ . Rinse your last time with distilled water. Then place the labeled beaker in the drying oven.
- c. Mass the dry nails, and then discard the nails.

Step 3

- a. Mass the beaker + dry copper. Discard the copper in the place designated by your teacher. Wash your beaker and let dry.

Data:

Mass 250 mL beaker	
Mass 250 mL beaker + copper (II) chloride	
Mass nails before reaction	
Mass nails after reaction	

Modeling Chemistry Nail Lab

M Walker



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