Name \_\_\_\_\_\_\_ Pd

# Unit 7: Balancing Chemical Reactions · Worksheet 2

Balance the following equations by inserting the proper coefficients.

- 1.  $_2$ \_SO<sub>2</sub> +  $_1$ \_O<sub>2</sub>  $\rightarrow$   $_2$ \_SO<sub>3</sub>
- 2.  $2 \text{ CH4} + 3 \text{ O2} \rightarrow 2 \text{ CO} + 4 \text{ H2O}$
- 2 P + 3 G2 → 2 PG3
- 2\_CO + 1\_O<sub>2</sub> → 2\_CO<sub>2</sub>
- 5.  $_2$ CH4 +  $_1$ O2  $\rightarrow$   $_2$ CH3OH
- 2\_Li + 1\_Br2 → 2\_LiBr
- 7.  $_2AI_2O_3 \rightarrow _4AI + _3O_2$
- 2 Na + 2 H<sub>2</sub>O → 2 NaOH + 1 H<sub>2</sub>
- 9.  $_{6}$ \_CO<sub>2</sub> +  $_{6}$ \_H<sub>2</sub>O  $\rightarrow$   $_{1}$ \_C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> +  $_{6}$ \_O<sub>2</sub>
- 10. \_1\_H2SO4 + \_2\_NaCl → \_2\_HCl + \_1\_Na2SO4

Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:

- 11. 3 H<sub>2</sub> + 1 SO<sub>2</sub> → 1 H<sub>2</sub>S + 2 H<sub>2</sub>O
- 12.  $_2$ CaCO3 +  $_2$ SO2 +  $_1$ O2  $\rightarrow$   $_2$ CaSO4 +  $_2$ CO2
- 2\_AgNO<sub>3</sub> + 1\_CaCl<sub>2</sub> → 2\_AgCl + 1\_Ca(NO<sub>3</sub>)<sub>2</sub>
- 14. 2 HCl + 1 Ba(OH)2 → 1 BaCl2 + 2 H2O
- 15. \_1\_H3PO4 + \_3\_NaOH → \_1\_Na3PO4 + \_3\_H2O
- 16. \_1\_Pb(NO3)2 + \_2\_KI → \_1\_PbI2 + \_2\_KNO3
- 17.  $3 \text{ O}_1\text{ O} + 2 \text{ NH}_3 \rightarrow 1 \text{ N}_2 + 3 \text{ O}_1 + 3 \text{ H}_2\text{ O}_1$
- 18.  $_1$ C<sub>2</sub>H<sub>5</sub>OH +  $_3$ C<sub>2</sub>  $\rightarrow$   $_2$ CO<sub>2</sub> +  $_3$ H<sub>2</sub>O
- 19.  $_2$ C2H6 +  $_3$ O2  $\rightarrow$   $_2$ CH3COOH +  $_2$ H2O
- 20.  $3NO_2 + 1H_2O \rightarrow 2HNO_3 + 1NO$

# **Modeling Chemistry Worksheet 1 Rearranging Atoms Answer Key**

Judd E. Hollander

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