

$$\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100\%$$

↳ from B-E-A chart

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation.
- Identify the given (with units) and what you want to find (with units).
- Show set up with units. Give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



1 mol H_2O : 18.02 g/mol

36 g $\frac{1\text{ mol H}_2\text{O}}{18.02\text{ g}} = 1.99\text{ mol H}_2\text{O}$

1 mol H_2 : 2.02 g

2 mol $\text{H}_2 = \boxed{4.04\text{ g H}_2}$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



1 mol $\text{NaCl} = 58.44\text{ g}$
 $581\text{ g NaCl} \times \frac{1\text{ mol NaCl}}{58.44\text{ g}} = 9.94\text{ mol NaCl}$

1 mol $\text{Na} = 22.99\text{ g}$

$9.94\text{ mol Na} \times 22.99\text{ g} = \boxed{228.52\text{ g Na}}$

Modeling Chemistry Unit 8 Test

M Lipman



Modeling Chemistry Unit 8 Test:

Modeling Chemistry Unit 8 Test Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Modeling Chemistry Unit 8 Test**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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