

MIXED GAS LAWS WORKSHEET

Key

Directions: Examine each question and then write the form of the gas law you plan to use to solve each question. Show which values you have, which values are missing and/or which values need to be calculated. Be careful to use standard units of volume (liters), temperature (Kelvins) & pressure (Atm or mm of Hg).

Note: 1 atm = 760 mm Hg

1. A gas occupies 3.5L at 2.5 mm Hg pressure. What is the volume at 10 mm Hg at the same temperature?

$$P_1 = 2.5 \text{ mmHg} \quad P_2 = 10 \text{ mmHg} \quad P_1 V_1 = P_2 V_2 \quad \frac{(2.5 \text{ mmHg})(3.5 \text{ L})}{10 \text{ mmHg}} = \frac{(10 \text{ mmHg}) V_2}{10 \text{ mmHg}}$$

$$V_1 = 3.5 \text{ L} \quad V_2 = ? \quad \boxed{0.875 \text{ L} = V_2}$$

2. A constant volume of oxygen is heated from 100°C to 185°C. The initial pressure is 4.1 atm. What is the final pressure?

$$P_1 = 4.1 \text{ atm} \quad \frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \frac{4.1 \text{ atm}}{373 \text{ K}} = \frac{P_2}{458 \text{ K}}$$

$$T_1 = 100^\circ\text{C} + 273 = 373 \text{ K} \quad P_2 = \frac{(4.1 \text{ atm})(458 \text{ K})}{373 \text{ K}} = \boxed{5.0 \text{ atm}}$$

$$P_2 = ? \quad T_2 = 185^\circ\text{C} + 273 = 458 \text{ K}$$

3. A sample of 25L of NH_3 gas at 10°C is heated at constant pressure until it fills a volume of 50L. What is the new temperature in °C?

$$V_1 = 25 \text{ L} \quad \frac{25 \text{ L}}{283 \text{ K}} = \frac{50 \text{ L}}{T_2} \quad T_2 = \frac{(50 \text{ L})(283)}{25 \text{ L}} = 566 \text{ K}$$

$$T_1 = 10^\circ\text{C} + 273 = 283 \text{ K} \quad ^\circ\text{C} = \text{K} - 273 = 566 - 273 = \boxed{293^\circ\text{C}}$$

$$V_2 = 50 \text{ L} \quad T_2 = ?$$

4. An unknown gas weighs 34g and occupies 6.7L at 2 atm and 245K. What is its molecular weight?

$$\text{Molar Mass} = \frac{g}{\text{mol}} \quad PV = nRT \quad (2 \text{ atm})(6.7 \text{ L}) = n (0.0821 \frac{\text{L atm}}{\text{mol K}})(245 \text{ K})$$

$$n = \frac{(2 \text{ atm})(6.7 \text{ L})}{(0.0821 \frac{\text{L atm}}{\text{mol K}})(245 \text{ K})} = 0.666 \text{ mol}$$

$$\text{Molar mass} = \frac{34 \text{ g}}{0.666 \text{ mol}} = \boxed{51 \text{ g/mol}}$$

5. An ideal gas occupies 400ml at 270 mm Hg and 65°C. If the pressure is changed to 1.4 atm and the temperature is increased to 100°C, what is the new volume?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \frac{(0.355 \text{ atm})(400 \text{ mL})}{318 \text{ K}} = \frac{(1.4 \text{ atm})(V_2)}{373 \text{ K}}$$

$$V_2 = \frac{(0.355 \text{ atm})(400 \text{ mL})(373 \text{ K})}{(318 \text{ K})(1.4 \text{ atm})} = \boxed{112 \text{ mL}}$$

$$\text{Unit: } \begin{cases} P_1 = 270 \text{ mmHg} \\ V_1 = 400 \text{ mL} \\ T_1 = 65^\circ\text{C} + 273 = 338 \text{ K} \\ P_2 = 1.4 \text{ atm} \\ V_2 = ? \\ T_2 = 100^\circ\text{C} + 273 = 373 \text{ K} \end{cases}$$

6. What is the volume of 23g of neon gas at 1°C and a pressure of 2 atm?

$$PV = nRT \quad 23 \text{ g Ne} \left(\frac{1 \text{ mol Ne}}{20.18 \text{ g Ne}} \right) = 1.14 \text{ mol Ne}$$

$$V = \frac{nRT}{P} = \frac{(1.14 \text{ mol})(0.0821 \frac{\text{L atm}}{\text{mol K}})(274 \text{ K})}{2 \text{ atm}} = \boxed{12.8 \text{ L}}$$

Mixed Gas Law Calculations Answers

SJ Ball



Mixed Gas Law Calculations Answers:

Decoding **Mixed Gas Law Calculations Answers**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Mixed Gas Law Calculations Answers**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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