

$$PV = nRT$$

Gas Stoichiometry Worksheet

$$PV = \frac{m}{M}RT \quad \frac{m}{V} = \frac{PM}{RT}$$

Name: Key

Date: _____

Period: _____

Volume - Volume

1. In the electrolysis of water, 75.0 mL of oxygen are produced. How many milliliters of hydrogen are produced? (Ans: 150. mL)



2. How many liters of hydrogen are required to react with excess nitrogen to produce 50.0 L ammonia (NH₃)? (Ans: 75.0 L)



3. Ammonia burns in oxygen to give nitrogen monoxide (NO) and water. If 20.0 L of ammonia are burned, how many liters of oxygen are required? (Ans: 25.0 L)

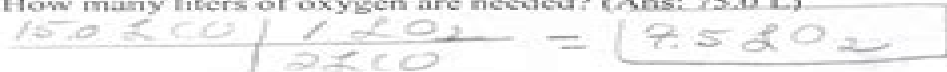


4. Carbon monoxide burns in oxygen and forms carbon dioxide.

- a. How many liters of carbon dioxide are produced from 15.0 L of CO? (Ans: 15.0 L)



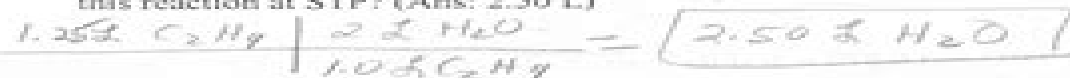
- b. How many liters of oxygen are needed? (Ans: 7.5 L)



5. If 400 mL of hydrogen and 400 mL of oxygen are mixed and ignited, what volume of oxygen remains uncombined? (Ans: 200 mL)



6. Ethylene burns in oxygen to form carbon dioxide and water vapor in the following reaction: $C_2H_4(g) + 3O_2(g) \rightarrow 2CO_2(g) + 2H_2O(g)$. How many liters of water can be formed if 1.25 liters of ethylene are consumed in this reaction at STP? (Ans: 2.50 L)



Mass - Volume



1. How many grams of sodium are needed to release 4.0 L of hydrogen from water? (Ans: 8.21 g)



2. When 130 g of zinc react with excess hydrochloric acid, what volume of hydrogen is produced? (Ans: 44.8 L)



Mixed Stoichiometry Problems Key

RJ Shavelson



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