

Stoichiometry Reflect #2

Learning Target... I can use stoichiometry to calculate amounts

1a. Balance this equation



1b. Determine how many moles of ammonium oxide was used if the student weighed out 55.2 g of it.

$$55.2 \text{ g (NH}_4\text{)}_2\text{O} \left(\frac{1 \text{ mol A}}{52.10 \text{ g}} \right) = \boxed{1.06 \text{ moles (NH}_4\text{)}_2\text{O}}$$

1c. Determine the amount of moles of ammonium nitride produced in this reaction.

$$1.06 \text{ moles A} \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) = \boxed{0.706 \text{ moles (NH}_4\text{)}_3\text{N}}$$

1d. Determine how many grams of ammonium nitride produced in this reaction.

$$0.706 \text{ moles} \left(\frac{108.11 \text{ g}}{1 \text{ mol B}} \right) = \boxed{48.1 \text{ g (NH}_4\text{)}_3\text{N}}$$

2. Using the following chemical reaction:

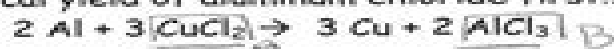


and, knowing that 61.6 g of copper was produced in this experiment, how many molecules of mercury nitrate was also produced?

$$61.6 \text{ g Cu} \left(\frac{1 \text{ mol A}}{63.55 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{1 \text{ mol A}} \right) \left(\frac{6.02 \times 10^{23} \text{ mc}}{1 \text{ mol B}} \right) = \boxed{1.17 \times 10^{24} \text{ mc HgNO}_3}$$

3. Challenge...

What is the percent yield of the following reaction if 12.8 g of copper (II) chloride was used in the lab, and 7.75 g of aluminum chloride was measured at the end of the experiment? (hint: you must calculate the theoretical yield of aluminum chloride first!!)



$$12.8 \text{ g A} \left(\frac{1 \text{ mol A}}{134.45 \text{ g}} \right) \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) \left(\frac{133.33 \text{ g}}{1 \text{ mol B}} \right) = 8.46 \text{ g AlCl}_3$$

theoretical

$$\% \text{ yield} = \frac{7.75 \text{ g}}{8.46 \text{ g}} = \boxed{91.4\%}$$

Show work on back

Mixed Stoichiometry Practice Answer Key

**Donald J. Wink, Sharon Fetzer-
Gislason, Sheila McNicholas**



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Table of Contents Mixed Stoichiometry Practice Answer Key

1. Understanding the eBook Mixed Stoichiometry Practice Answer Key
 - The Rise of Digital Reading Mixed Stoichiometry Practice Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Stoichiometry Practice Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mixed Stoichiometry Practice Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Stoichiometry Practice Answer Key
 - Personalized Recommendations
 - Mixed Stoichiometry Practice Answer Key User Reviews and Ratings

- Mixed Stoichiometry Practice Answer Key and Bestseller Lists
- 5. Accessing Mixed Stoichiometry Practice Answer Key Free and Paid eBooks
 - Mixed Stoichiometry Practice Answer Key Public Domain eBooks
 - Mixed Stoichiometry Practice Answer Key eBook Subscription Services
 - Mixed Stoichiometry Practice Answer Key Budget-Friendly Options
- 6. Navigating Mixed Stoichiometry Practice Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Mixed Stoichiometry Practice Answer Key Compatibility with Devices
 - Mixed Stoichiometry Practice Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mixed Stoichiometry Practice Answer Key
 - Highlighting and Note-Taking Mixed Stoichiometry Practice Answer Key
 - Interactive Elements Mixed Stoichiometry Practice Answer Key
- 8. Staying Engaged with Mixed Stoichiometry Practice Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mixed Stoichiometry Practice Answer Key
- 9. Balancing eBooks and Physical Books Mixed Stoichiometry Practice Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mixed Stoichiometry Practice Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mixed Stoichiometry Practice Answer Key
 - Setting Reading Goals Mixed Stoichiometry Practice Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mixed Stoichiometry Practice Answer Key
 - Fact-Checking eBook Content of Mixed Stoichiometry Practice Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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