

Lab 1- What is around us?

2/1/2024

Focus Questions

1) What types of compounds do you know? Give a few examples.

- Two examples of compounds are Ionic and covalent bonds.

2) When naming different types of compounds do you need to follow different rules? Explain using examples.

- For naming covalent bonds first write the names of both elements. Then changes the ending of the second one to ide and add the necessary prefixes.
 - $\text{SiCl}_4 \rightarrow$ Silicon Tetrachloride
- For ionic compounds, write the metal's name first and then the non-metal. Change non-metals ending to ide. Add roman numerals if necessary.
 - $\text{MgBr}_2 \rightarrow$ Magnesium Bromide
 - $\text{CuBr} \rightarrow$ Copper (I) Bromide

3) Does it matter in which order you combine your elements when writing chemical formulas for ionic compounds?

- Yes it does. The cation is typically written first which is the atom which is losing an electron. This means the positive ion is written first and then the negative ion is. The polyatomic ion is always written last.

4) Which laws govern the formation of compounds?

- The law of mass action.

Post-Lab Assessment Questions (20 points):

1. (6 pts) For the following pairs of ions, write the formula of the compound that you would expect them to form:

- a. barium and cyanide $\text{Ba}(\text{CN})_2$
- b. cobalt(III) and sulfur Co_2S_3
- c. iron(III) and sulfate ~~FeSO₄~~ $\text{Fe}_2(\text{SO}_4)_3$

2. (6 pts) Platinum is a transition metal and forms two different ions, Pt^{2+} and Pt^{4+} . Write the formulas for the compounds for each platinum ion with:

- a. chloride ions PtCl_2 , PtCl_4
- b. carbonate ions PtCO_3 , $\text{Pt}(\text{CO}_3)_2$
- c. nitrate ions $\text{Pt}(\text{NO}_3)_2$, $\text{Pt}(\text{NO}_3)_4$

3. (4 pts) Name the following compounds:

- a. N_2O_5 Dinitrogen Pentaoxide
- b. AlCl_3 ~~Aluminum Chloride~~ Aluminium Chloride
- c. K_3PO_4 Tri potassium Phosphate
- d. MnO_2 Manganese (IV) Oxide

4. (4 pts) Write the formulas for the following compounds:

- a. Sulfur tetrachloride SCl_4
- b. Bromine dioxide BrO_2
- c. Nickel(II) nitrate $\text{Ni}(\text{NO}_3)_2$
- d. Calcium sulfide CaS

Hunter Lab Manual Chem 106 Lab

CL Gary



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Table of Contents Hunter Lab Manual Chem 106 Lab

1. Understanding the eBook Hunter Lab Manual Chem 106 Lab
 - The Rise of Digital Reading Hunter Lab Manual Chem 106 Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Hunter Lab Manual Chem 106 Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hunter Lab Manual Chem 106 Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hunter Lab Manual Chem 106 Lab
 - Personalized Recommendations

- Hunter Lab Manual Chem 106 Lab User Reviews and Ratings
- Hunter Lab Manual Chem 106 Lab and Bestseller Lists
- 5. Accessing Hunter Lab Manual Chem 106 Lab Free and Paid eBooks
 - Hunter Lab Manual Chem 106 Lab Public Domain eBooks
 - Hunter Lab Manual Chem 106 Lab eBook Subscription Services
 - Hunter Lab Manual Chem 106 Lab Budget-Friendly Options
- 6. Navigating Hunter Lab Manual Chem 106 Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Hunter Lab Manual Chem 106 Lab Compatibility with Devices
 - Hunter Lab Manual Chem 106 Lab Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hunter Lab Manual Chem 106 Lab
 - Highlighting and Note-Taking Hunter Lab Manual Chem 106 Lab
 - Interactive Elements Hunter Lab Manual Chem 106 Lab
- 8. Staying Engaged with Hunter Lab Manual Chem 106 Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hunter Lab Manual Chem 106 Lab
- 9. Balancing eBooks and Physical Books Hunter Lab Manual Chem 106 Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hunter Lab Manual Chem 106 Lab
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hunter Lab Manual Chem 106 Lab
 - Setting Reading Goals Hunter Lab Manual Chem 106 Lab
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hunter Lab Manual Chem 106 Lab
 - Fact-Checking eBook Content of Hunter Lab Manual Chem 106 Lab

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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