

Solid State

Types of solids

Crystalline solid	Amorphous solids
- Definite & regular arrangement - True solids - Long ordered arrangement of particles - These are anisotropic - Sharp M.P - Clean cleavage cut - Ex - NaCl, KCl	- Doesn't have regular arrange - Pseudo solids or super cooled liq - Short ordered arrangement - These are isotropic - Melt over certain range of temp - Irregular cleavage cut - Ex - glass, rubber

Isomorphism - Two or more substance having same crystal structure
 - same atomic ratio
 Ex NaF & MgO - 1:1
 NaNO₃ & CaCO₃ - 1:1:3

Polymorphism - A single substance that exists in two or more forms of crystalline structure
 Ex Calcite & aragonite are forms of calcium carbonate
 Ex - α-quartz, β-quartz & cristobalite - these forms of silica
 - Polymorphism occurring in elements are called allotropy
 Ex Carbon - diamond, graphite, fullerene.

Classification of crystalline solids

1) **Ionic** - Electrostatic force of attraction
 Ion - Ion force
 - Brittle, hard, high melting point
 - non conductors of electricity "when dissolved or melted in water."
 Ex - NaCl, KCl, MgCl₂

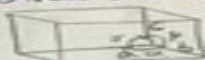
2) **Molecular** -
 - Dispersion / Dipole-Dipole / H-Bond
 - Soft, low melting point
 - non conductors & good insulators
 Ex H₂O, Br₂, CO₂, CH₄

3) **Covalent** -
 - Covalent Bond
 - Hard, high melting point, B.P
 - Most incompressible
 - Poor conductor
 Ex - C-diamond, SiO₂, BN, Carborandum

4) **Metallic** -
 - Metallic Bond
 - Variable hardness, M.P, conductivity
 - They are malleable, ductile & good conductors
 Ex - Na, Zn, Cu, Fe

Crystal Lattice - or space lattice - Lattice is a geometrical arrangement of points in 3D form such constituent particles that are attached to the lattice points are crystal lattice.

Unit Cell - The smallest repeating structural unit of a crystalline solid



Ex of unit cell.

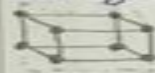
1) Simple or primitive unit cell

- Particles are present at corners only
 $8 \times \frac{1}{8} = 1 \text{ atom}$

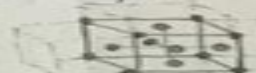
2) **Face centered** - The particles are present at the corners & also at the center
 $8 \times \frac{1}{8} + 6 \times \frac{1}{2} = 1 + 3 = 4 \text{ atoms}$

3) **Body centered** - The particles are present at the corners as well as at the centre of the unit cell
 $8 \times \frac{1}{8} + 1 = 1 + 1 = 2 \text{ atoms}$

Note -
 • An atom at the corner is shared by eight unit cell - $\frac{1}{8}$
 • An atom at the face is shared by two cell - $\frac{1}{2}$
 • An atom at the body centre of a unit cell belongs entirely to it - 1



$8 \times \frac{1}{8} = 1$
 1 atom
 SC - 1



$8 \times \frac{1}{8} = 1$
 $6 \times \frac{1}{2} = 3$
 $1 + 3 = 4 \text{ atoms}$
 FCC - 4



$8 \times \frac{1}{8} = 1$
 $1 \times 1 = 1$
 $1 + 1 = 2 \text{ atoms}$
 BCC - 2

Hsc Chemistry Solid State Notes

AW Rasmussen



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