

## Experiment 1: Charles's Law

Table 1: Temperature and Volume Data

Temperature Conditions	Temperature (°C)	Temperature (K)	Syringe Volume (mL)
Room Temperature	23 C	296.15 K	5 ml
Hot Water	45 C	318.15 K	1 ml
Ice Water	5 C	278.15 K	0.3 ml

### Questions:

1. Create a graph of temperature and volume data. Place temperature on the x-axis (in Kelvin) and volume (mL) on the y-axis. Insert your graph below.
2. What happened to the volume of gas when the syringe was exposed to various temperature conditions? Using the concepts explored in the Introduction, describe why this occurred, keeping in mind the definition of temperature.
3. Draw a straight line of best fit through your data points on your graph in question 1, extrapolating the line until it intersects the (negative) x-axis. Why can you assume a linear relationship (a straight-lined slope)?
4. At what temperature does your line intersect the x-axis? What volume corresponds to this temperature?

# Escience Lab Ideal Gas Law

**S Ashworth**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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where  $V$  is the volume of the container shared between each particle not the velocity of a particle  $N$  is the total number of particles and  $k_B$  is Boltzmann's constant given by the universal gas constant  $R$  and Avogadro's constant  $N_A$  Extra parameters are introduced  $a$  is a measure for the attraction between the particles and  $b$  is the average volume excluded from  $V$  by a particle The equation can be cast into the better known form where  $a$  is a measure of the attraction between the particles  $b$  is the volume excluded by a mole of particles A careful distinction

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