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web sep 13 2023 the only difference between molality and molarity is in the denominator molality is measured in kilograms of solvent whereas molarity is measured in liters of solution when studying the properties of solutions related to vapour pressure and temperature changes concentrations expressed in molality are used

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web molarity has units of $\frac{\text{mol}}{\text{L}}$ text mol text liter litermol which can be abbreviated as molar or text m m pronounced molar the molar concentration of the solute is sometimes abbreviated by putting square brackets around the chemical formula of

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web molarity is the number of moles of a substance per litre of solution also known as molar concentration a capital m signifies solutions labelled with molar concentration a 1 0 m solution contains 1 mole of solvent per litre of solution molality is the number of solvent moles per kilogram

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web 1 molarity mv mass molar mass x 0 4500 l 0 825 g 141 9579 g mol x 0 0129 m 2 molality 0 825 g 141 9579 g mol 0 00581158 mol 0 00581158 mol 0 4500 kg 0 0129 m 3 mole fraction na 2 hpo 4 0 825 g 141 9579 g mol 0 00581158 mol h 2 o 450 0 g 18 015 g mol 24 97918401 mol

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web apr 7 2014 molarity moles of solute liters of solution 8 4 2 2 a first convert 250 ml to liters 250 1000 0 25 then calculate molarity 5 moles 0 25 liters 20 m 3 c a solution with molarity 2 requires 2 m of n a oh per liter so 4 x 2 8 m 4 a a solution of molarity 1 5 m requires 1 5 mol of na to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture mmix can be calculated using the following formula $m_{\text{mix}} = \frac{m_1 + m_2 + \dots + m_n}{V_1 + V_2 + \dots + V_n}$ 1 2 x 480 1 2 x 520 480 520 720 624 1000 1 344 m q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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molarity of a solution containing 325 g of nacl dissolved in 750 ml of solution 3 140 g of kcl is dissolved in 600 ml of water what is the molarity 4 724 4 g of ammonium phosphate in 4500 ml of alcohol

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web determine the molarity for each of the following solutions 0 444 mol of cocl_2 in 0 654 l of solution 98 0 g of phosphoric acid h_3po_4 in 1 00 l of solution 0 2074 g of calcium hydroxide $\text{ca}(\text{oh})_2$ in 40 00 ml of solution 10 5 kg

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