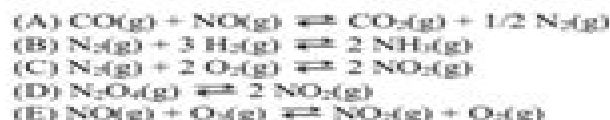


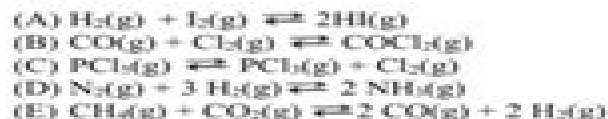


When the substances in the equation above are at equilibrium at pressure P and temperature T, the equilibrium can be shifted to favor the products by

- (A) increasing the pressure by means of a moving piston at constant T
 (B) increasing the pressure by adding an inert gas such as nitrogen
 (C) decreasing the temperature
 (D) allowing some gases to escape at constant P and T
 (E) adding a catalyst
2. In which of the following systems would the number of moles of the substances present at equilibrium NOT be shifted by a change in the volume of the system at constant temperature?



3. At constant temperature, a change in volume will NOT affect the moles of substances present in which of the following?



4. $\text{PCl}_5\text{(g)} + \text{Cl}_2\text{(g)} \rightleftharpoons \text{PCl}_3\text{(g)} + \text{energy}$
 Some PCl_5 and Cl_2 are mixed in a container at 200 °C and the system reaches equilibrium according to the equation above. Which of the following causes an increase in the number of moles of PCl_5 present at equilibrium?

- I. Decreasing the volume of the container
 II. Raising the temperature
 III. Adding a mole of He gas at constant volume

- (A) I only
 (B) II only
 (C) I and III only
 (D) II and III only
 (E) I, II, and III



When 0.40 mole of SO_2 and 0.60 mole of O_2 are placed in an evacuated 1.00-liter flask, the reaction represented above occurs. After the reactants and the product reach equilibrium and the initial temperature is restored, the flask is found to contain 0.30 mole of SO_3 . Based on these results, the expression for the equilibrium constant, K_c , of the reaction is

- (A) $(0.30)^2 / [(0.45)(0.10)^2]$
 (B) $(0.30)^2 / [(0.60)(0.40)^2]$
 (C) $(2 \times 0.30) / [(0.45)(2 \times 0.10)]$
 (D) $(0.30) / [(0.45)(0.10)]$
 (E) $(0.30) / [(0.60)(0.40)]$

6. How many moles of NaF must be dissolved in 1.00 liter of a saturated solution of PbF_2 at 25 °C to reduce the $[\text{Pb}^{2+}]$ to 1×10^{-8} molar?

(K_{sp} of PbF_2 at 25 °C = 4.0×10^{-8})

- (A) 0.020 mole
 (B) 0.040 mole
 (C) 0.10 mole
 (D) 0.20 mole
 (E) 0.40 mole

7. The solubility of CuI is 2×10^{-8} molar. What is the solubility product constant, K_{sp} , for CuI ?

- (A) 1.4×10^{-17}
 (B) 2×10^{-16}
 (C) 4×10^{-12}
 (D) 2×10^{-12}
 (E) 8×10^{-18}

8. What is the molar solubility in water of Ag_2CrO_4 ? (The K_{sp} for Ag_2CrO_4 is 8×10^{-12} .)

- (A) $8 \times 10^{-12} \text{ M}$
 (B) $2 \times 10^{-12} \text{ M}$
 (C) $(4 \times 10^{-12} \text{ M})^{1/2}$
 (D) $(4 \times 10^{-12} \text{ M})^{1/3}$
 (E) $(2 \times 10^{-12} \text{ M})^{1/3}$

9. A student wishes to reduce the zinc ion concentration in a saturated zinc iodate solution to $1 \times 10^{-6} \text{ M}$. How many moles of solid KIO_3 must be added to 1.00 L of solution? ($K_{sp} \text{ Zn(IO}_3)_2 = 4 \times 10^{-6}$ at 25°C)

- (A) 1 mol
 (B) 0.5 mol
 (C) 0.0001 mol
 (D) 4 mol
 (E) 2 mol

Equilibrium Test 2 Multiple Choice Answers

Arshad Iqbal



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Table of Contents Equilibrium Test 2 Multiple Choice Answers

1. Understanding the eBook Equilibrium Test 2 Multiple Choice Answers
 - The Rise of Digital Reading Equilibrium Test 2 Multiple Choice Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Equilibrium Test 2 Multiple Choice Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Equilibrium Test 2 Multiple Choice Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Equilibrium Test 2 Multiple Choice Answers
 - Personalized Recommendations
 - Equilibrium Test 2 Multiple Choice Answers User Reviews and Ratings
 - Equilibrium Test 2 Multiple Choice Answers and Bestseller Lists
5. Accessing Equilibrium Test 2 Multiple Choice Answers Free and Paid eBooks
 - Equilibrium Test 2 Multiple Choice Answers Public Domain eBooks
 - Equilibrium Test 2 Multiple Choice Answers eBook Subscription Services

- Equilibrium Test 2 Multiple Choice Answers Budget-Friendly Options
- 6. Navigating Equilibrium Test 2 Multiple Choice Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Equilibrium Test 2 Multiple Choice Answers Compatibility with Devices
 - Equilibrium Test 2 Multiple Choice Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Equilibrium Test 2 Multiple Choice Answers
 - Highlighting and Note-Taking Equilibrium Test 2 Multiple Choice Answers
 - Interactive Elements Equilibrium Test 2 Multiple Choice Answers
- 8. Staying Engaged with Equilibrium Test 2 Multiple Choice Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Equilibrium Test 2 Multiple Choice Answers
- 9. Balancing eBooks and Physical Books Equilibrium Test 2 Multiple Choice Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Equilibrium Test 2 Multiple Choice Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Equilibrium Test 2 Multiple Choice Answers
 - Setting Reading Goals Equilibrium Test 2 Multiple Choice Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Equilibrium Test 2 Multiple Choice Answers
 - Fact-Checking eBook Content of Equilibrium Test 2 Multiple Choice Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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