

Empirical and Molecular Formula Practice Solutions

1. A compound is found to contain 63.52 % iron and 36.48% sulfur. Find its empirical formula.

$$63.52 \text{ g Fe} \times \frac{1 \text{ mol}}{55.85 \text{ g}} = 1.137 \text{ mol Fe}$$

$$36.48 \text{ g S} \times \frac{1 \text{ mol}}{32.06 \text{ g}} = 1.138 \text{ mol S}$$

$$\frac{1.137 \text{ mol Fe}}{1.137} : \frac{1.138 \text{ mol S}}{1.137} = 1 \text{ mol Fe: } 1.001 \text{ mol S}$$

empirical formula = FeS

2. A compound is found to contain 26.56% potassium, 35.41% chromium, and the remainder oxygen. Find its empirical formula.

$$26.56 \text{ g K} \times \frac{1 \text{ mol}}{39.10 \text{ g}} = 0.6793 \text{ mol K}$$

$$36.48 \text{ g Cr} \times \frac{1 \text{ mol}}{52.00 \text{ g}} = 0.7015 \text{ mol Cr}$$

$$36.96 \text{ g O} \times \frac{1 \text{ mol}}{16.00 \text{ g}} = 2.310 \text{ mol O}$$

$$\frac{0.6793 \text{ mol K}}{0.6793} : \frac{0.7015 \text{ mol Cr}}{0.6793} : \frac{2.310 \text{ mol O}}{0.6793} = 1 \text{ mol K: } 1.033 \text{ mol Cr: } 3.401 \text{ mol O}$$

$$\rightarrow 2 \text{ mol K: } 2.066 \text{ mol Cr: } 6.802 \text{ mol O}$$

empirical formula = $\text{K}_2\text{Cr}_2\text{O}_7$

3. A 60.00 g sample of tetraethyllead, a gasoline additive, is found to contain 38.43 g lead, 17.83 g carbon, and 3.74 g hydrogen. Find its empirical formula.

$$38.43 \text{ g Pb} \times \frac{1 \text{ mol}}{207.20 \text{ g}} = 0.1855 \text{ mol Pb}$$

$$17.83 \text{ g C} \times \frac{1 \text{ mol}}{12.01 \text{ g}} = 1.485 \text{ mol C}$$

$$3.74 \text{ g H} \times \frac{1 \text{ mol}}{1.01 \text{ g}} = 3.78 \text{ mol H}$$

$$\frac{0.1855 \text{ mol Pb}}{0.1855} : \frac{1.485 \text{ mol C}}{0.1855} : \frac{3.78 \text{ mol H}}{0.1855} = 1 \text{ mol Pb: } 8.005 \text{ mol C: } 20.4 \text{ mol H}$$

empirical formula = $\text{PbC}_8\text{H}_{20}$

Empirical And Molecular Formula Calculation Practice

Neil D. Jespersen, Pamela Kerrigan



Empirical And Molecular Formula Calculation Practice:

Decoding **Empirical And Molecular Formula Calculation Practice**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Empirical And Molecular Formula Calculation Practice**," a mesmerizing literary creation penned with a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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