

Activity A:

Subatomic particles

Get the Gizmo ready:

- Use the arrows to create an atom with two protons, two neutrons, and two electrons.
- Turn on **Show element name**.



Question: What are the properties of protons, neutrons, and electrons?

1. **Observe:** Turn on **Show element symbol** and **Element notation**. Three numbers surround the element symbol: the **mass number** (A), electrical charge (no number is displayed if the atom is neutral), and the **atomic number** (Z).
2. **Investigate:** Watch how the numbers change as you add or remove particles.



- A. Which number is equal to the number of protons in the atom?
- B. How can you calculate the number of neutrons (N) in an atom?
- C. Which particle (proton, neutron, or electron) has a
Positive charge?
Negative charge?

neutrons

by the number of protons

electron

proton

neutron

3. **Observe:** Turn on **Isotope notation**. Isotope notation shows the mass number and atomic number. The mass number is the sum of protons and neutrons. The atomic number is the number of protons. The number of electrons is equal to the number of protons in a neutral atom.

- A. What are the mass numbers of carbon?
- B. What are the mass numbers of oxygen?
- C. Carbon has two isotopes. What are they?

12 and 13

16 and 17

12

4. **Observe:** Turn on **Isotope notation** and **Element notation**.

- A. How many electrons are in a neutral atom of lithium?
- B. How many electrons are in an atom of hydrogen?
- C. What is the mass number of an atom with 3 protons and 3 neutrons?
- D. What is the mass number of an atom with 11 protons and 11 neutrons?
- E. How many electrons are in hydrogen?

3

1

6

11

1

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