# MODELS OF THE ATOM

# **Section Review**

#### Objectives

- · Identify inadequacies in the Rutherford atomic model
- · Identify the new assumption in the Bohr model of the atom
- Describe the energies and positions of electrons according to the quantum mechanical model
- Describe how the shapes of orbitals at different sublevels vary

#### Vocabulary

- · emergy levels
- quantum

- · quantum mechanical model
- · atomic orbital

#### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The chemical properties of atoms, ions, and molecules

are related to the arrangement of the 1 within them.

2.

in \_\_\_6\_\_ paths. The \_\_\_7\_\_ model is the modern description
of the electrons in atoms. This model estimates the \_\_\_\_8\_\_ of finding an
electron within a certain volume of space surrounding the nucleus.

#### Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- Electrons must have a certain minimum amount of energy called a quantum in order to move from one energy level to the next higher energy level.
- \_\_\_\_\_\_10. The electron probability clouds for atomic orbitals are spherical in shape.

# **Electrons In Atoms Workbook Answers**

**Jerry Jenkins** 

**Electrons In Atoms Workbook Answers:** 

#### **Electrons In Atoms Workbook Answers** Book Review: Unveiling the Power of Words

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