

## Section 8.5

### Objectives

- **Describe** how electronegativity is used to determine bond type.
- **Compare and contrast** polar and nonpolar covalent bonds and polar and nonpolar molecules.
- **Generalize** about the characteristics of covalently bonded compounds.

## Review Vocabulary

**electronegativity:** the relative ability of an atom to attract electrons in a chemical bond

### New Vocabulary

polar covalent bond

## Electronegativity and Polarity

**MAIN Idea** A chemical bond's character is related to each atom's attraction for the electrons in the bond.

**Real-World Reading Link** The stronger you are, the more easily you can do pull-ups. Just as people have different abilities for doing pull-ups, atoms in chemical bonds have different abilities to attract (pull) electrons.

## Electron Affinity, Electronegativity, and Bond Character

The type of bond formed during a reaction is related to each atom's attraction for electrons. Electron affinity is a measure of the tendency of an atom to accept an electron. Excluding noble gases, electron affinity increases with increasing atomic number within a period and decreases with increasing atomic number within a group. The scale of electronegativities—shown in **Figure 8.20**—allows chemists to evaluate the electron affinity of specific atoms in a compound. Recall from Chapter 6 that electronegativity indicates the relative ability of an atom to attract electrons in a chemical bond. Note that electronegativity values were assigned, whereas electron affinity values were measured.

**Electronegativity** The version of the periodic table of the elements shown in **Figure 8.20** lists electronegativity values. Note that fluorine has the greatest electronegativity value (3.98), while francium has the least (0.7). Because noble gases do not generally form compounds, individual electronegativity values for helium, neon, and argon are not listed. However, larger noble gases, such as xenon, sometimes bond with highly electronegative atoms, such as fluorine.

**Figure 8.20** Electronegativity values are derived by comparing an atom's attraction for shared electrons to that of a fluorine's atom attraction for shared electrons. Note that the electronegativity values for the lanthanide and actinide series, which are not shown, range from 1.12 to 1.7.

1 <b>H</b> 2.20																	5 <b>B</b> 2.04	6 <b>C</b> 2.55	7 <b>N</b> 3.04	8 <b>O</b> 3.44	9 <b>F</b> 3.98								
3 <b>Li</b> 0.98	4 <b>Be</b> 1.57																	13 <b>Al</b> 1.61	14 <b>Si</b> 1.90	15 <b>P</b> 2.19	16 <b>S</b> 2.58	17 <b>Cl</b> 3.16							
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19 <b>K</b> 0.82	20 <b>Ca</b> 1.00	21 <b>Sc</b> 1.36	22 <b>Ti</b> 1.54	23 <b>V</b> 1.63	24 <b>Cr</b> 1.66	25 <b>Mn</b> 1.55	26 <b>Fe</b> 1.83	27 <b>Co</b> 1.88	28 <b>Ni</b> 1.91	29 <b>Cu</b> 1.90	30 <b>Zn</b> 1.65	36 <b>Kr</b> 2.20	37 <b>Rb</b> 0.82	38 <b>Sr</b> 0.95	39 <b>Y</b> 1.27	40 <b>Zr</b> 1.33	41 <b>Nb</b> 1.6	42 <b>Mo</b> 2.16	43 <b>Tc</b> 2.10	44 <b>Ru</b> 2.2	45 <b>Rh</b> 2.28	46 <b>Pd</b> 2.20	47 <b>Ag</b> 1.93	48 <b>Cd</b> 1.69	49 <b>In</b> 1.78	50 <b>Sn</b> 1.96	51 <b>Sb</b> 2.05	52 <b>Te</b> 2.1	53 <b>I</b> 2.66
55 <b>Cs</b> 0.79	56 <b>Ba</b> 0.89	57 <b>La</b> 1.10	72 <b>Hf</b> 1.3	73 <b>Ta</b> 1.5	74 <b>W</b> 1.7	75 <b>Re</b> 1.9	76 <b>Os</b> 2.2	77 <b>Ir</b> 2.2	78 <b>Pt</b> 2.2	79 <b>Au</b> 2.4	80 <b>Hg</b> 1.9	83 <b>Bi</b> 1.9	84 <b>Po</b> 2.0	85 <b>At</b> 2.2															
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**L Reisser**

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