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temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ J/g} \cdot \text{C} \times 110 \text{ g} \times 26.6 \text{ C} = 1220 \text{ J}$
 550 J $\Delta h_{\text{q water}} = 550 \text{ J}$

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 for water $4.184 \text{ J/g} \cdot \text{C}$ 1 what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased
 by 8.1 C when 207 J of heat is added 2

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 $14.7 \text{ J/g} \cdot \text{C}$ $\Delta H_{\text{rxn}} = q_{\text{rxn}} / n$ $q_{\text{calorimeter}} = m \cdot c \cdot \Delta T$

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web 6 a 218 g sample of steam at 121°C is cooled to ice at 14°C find the change in heat content of the system 7 in going from ice at 34°C to steam at 138°C a sample of water absorbs 1.41×10^5 J find the mass of the sample answers 1 2 68 x 104 J 2 6 23 x 104 J 3 1 71 x 104 J 4 1 11 x 105 J 5 2 11 x 105 J 6

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