

Hewlett-Packard Financial Performance Assessment

Exercise 10.17 has three similar following problems. Use all appropriate and proper scientific units. Use scientific notation for answers (e.g., 1.23×10^4).

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_{18}\text{O} + 15\text{O}_2 \rightarrow 10\text{CO}_2 + 9\text{H}_2\text{O}$.
 Element Data: $\text{C}_{10}\text{H}_{18}\text{O} = 1.23 \text{ kJ/mol}$ $\text{H}_2\text{O} = -285.8 \text{ kJ/mol}$
 and: $\text{CO}_2 = -393.5 \text{ kJ/mol}$ $\text{C}_{10}\text{H}_{18}\text{O} = 1.23 \text{ kJ/mol}$

a. $\Delta H_{\text{rxn}} = 1.23 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = 1.23 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = 1.23 \text{ kJ/mol}$
 b. $\Delta H_{\text{rxn}} = 1.23 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = 1.23 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = 1.23 \text{ kJ/mol}$

c. $\Delta H_{\text{rxn}} = 1.23 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = 1.23 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = 1.23 \text{ kJ/mol}$

2. Find standard enthalpy change for the formation of aluminum oxide and aluminum hydroxide.
 $\text{Al} + \frac{3}{2}\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$
 $\text{Al} + 3\text{H}_2\text{O} \rightarrow \text{Al}(\text{OH})_3$ $\Delta H_{\text{rxn}} = -1046.7 \text{ kJ/mol}$
 Standard enthalpy of formation: $\text{H}_2\text{O} = -285.8 \text{ kJ/mol}$ $\text{Al}(\text{OH})_3 = -1046.7 \text{ kJ/mol}$

a. $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$
 b. $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$

c. $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$

d. $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -1675.7 \text{ kJ/mol}$

3. Find Gibbs free energy change for the reaction of carbon monoxide and hydrogen, which can give for synthesis: $\text{CO} + 3\text{H}_2 \rightarrow \text{CH}_4 + \text{H}_2\text{O}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$

Calculate the standard enthalpy change for this reaction from the following standard enthalpies:

$\text{CO} = -110.5 \text{ kJ/mol}$ $\text{H}_2\text{O} = -285.8 \text{ kJ/mol}$ $\text{CH}_4 = -74.8 \text{ kJ/mol}$
 $\text{CO}_2 = -393.5 \text{ kJ/mol}$ $\text{H}_2\text{O} = -285.8 \text{ kJ/mol}$ $\text{CH}_4 = -74.8 \text{ kJ/mol}$

a. $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$
 b. $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$
 c. $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$

d. $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$
 e. $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$

f. $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$ $\Delta H_{\text{rxn}} = -131.3 \text{ kJ/mol}$

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Léo Pomar



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields **Algebraic and Discrete Mathematical Methods for Modern Biology** Raina Robeva, 2015-05-09 Written by experts in both mathematics and biology Algebraic and Discrete Mathematical Methods for

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Theory And Problems For Chemistry Olympiad: Challenging Concepts In Chemistry Zhihan Nan, Sheng Zhang, 2019-11-19 This study guide for the Chemistry Olympiad contains summarized concepts and examples in all areas of chemistry The chapters are arranged in a logical manner and establishes connections between concepts Undergraduate chemistry concepts are explained clearly every equation in physical chemistry is derived and justified while every organic reaction has its reaction mechanism shown and explained without assuming that readers have university level background in the subject The book also contains original Chemistry Olympiad sample problems that readers may use to test their knowledge This is a first book of its kind written by Nan Zhihan International Chemistry Olympiad IChO gold medallist and winner of the International Union of Pure and Applied Chemistry IUPAC Prize for achieving the highest score in the experimental exam and experienced Chemistry Olympiad trainer Dr Zhang Sheng who has served as head mentor of Singapore IChO team for many years It builds on the experience of both a participant and trainer to help any aspiring Chemistry Olympiad student understand the challenging concepts in chemistry Chemistry Workbook For Dummies with Online Practice Chris Hren, Peter J. Mikulecky, 2017-04-17 Take the confusion out of chemistry with hundreds of practice problems Chemistry Workbook For Dummies is your ultimate companion for introductory chemistry at the high school or college level Packed with hundreds of practice problems this workbook gives you the practice you need to internalize the essential concepts that form the foundations of chemistry From matter and molecules to moles and measurements these problems cover the full spectrum of topics you ll see in class and each section includes key concept review and full explanations for every problem to quickly get

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R.L.,2017 This textbook has been designed to meet the needs of B Sc Second Semester students of Chemistry as per the UGC Choice Based Credit System CBCS With its traditional approach to the subject this textbook lucidly explains principles of chemistry Important topics such as chemical energetics chemical ionic equilibrium aromatic hydrocarbons alkyl aryl halides alcohols phenols ethers aldehydes and ketones are aptly discussed to give an overview of physical and organic chemistry Laboratory work has also been included to help students achieve solid conceptual understanding and learn experimental procedures **Oswaal NCERT Exemplar (Problems - Solutions) Class 11 Physics, Chemistry and Mathematics (Set of 3 Books) For 2024 Exam** Oswaal Editorial Board,2023-10-28 Description of the product Chapter wise and Topic wise

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 questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry
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