

CHEMISTRY I EQUATION SHEET

Solubility Rules

Solubility Rules			Form a Compound which is	
Negative Ion Plus Positive Ion				
Any neg. ion + group 1a ions Soluble				
Any neg. ion + Ammonium ion Soluble				
Nitrate + any positive ion Soluble				
Acetate	+	any positive ion except Ag ⁺ or Hg ²⁺	Soluble	
bromide or iodide + Any other positive ion	+	Ag ⁺ , Pb ²⁺ , Hg ²⁺ or Cu ⁺	Not Soluble	
Sulfate	+	Ca ²⁺ , Sr ²⁺ , Ba ²⁺ , Ra ²⁺ , Ag ⁺ or Pb ²⁺	Not Soluble	
	+	Any other positive ion	Soluble	
Sulfide	+	group 1a ions or ammonium	Soluble	
	+	Be ²⁺ , Mg ²⁺ , Ca ²⁺ , Sr ²⁺ , Ba ²⁺ or Ra ²⁺	Soluble	
	+	Any other positive ion	Not Soluble	
Hydroxide	+	group 1a ions or ammonium	Soluble	
	+	Any other positive ion	Not Soluble	
Carbonate, Chromate, Phosphate, or Sulfite + Any other positive ion	+	group 1a ions or ammonium	Soluble	
	+	Any other positive ion	Not Soluble	

Equations

D = m/v	c = specific heat
Molarity = $\frac{\text{mol}}{\text{V}}$	d = density
$P_1 V_1 = P_2 V_2$	I = initial amount
$\frac{T_1}{T_2}$	F = final amount
Q = mc ΔT	m = mass
nRTPV =	M = molarity
F = I (1/2) *	n = number of moles & number of half lives
n = Total Time	P = pressure
Half Life Time	R = ideal gas constant
	Q = heat energy
	T = temperature
	V = volume
	X = amount of matter left

Constants

	1 atm = 101.3 kPa Pressure
	1 atm = 760 mm Hg
	R = 8.314 L x kPa / mol°C Universal
Constant or 0.0821 L x atm / mol°C	
Avogadro's number	6.02 X
Specific Heats (J / g°C)	
	Water = 4.180
	Aluminum = 0.895
	Copper = 0.387
	Iron = 0.448

Activity Series in Decreasing Activity

Metals
Lithium
Rubidium
Potassium
Calcium
Sodium
Magnesium
Aluminum
Manganese
Zinc
Iron
Nickel
Tin
Lead
Hydrogen
Copper
Silver
Platinum
Gold
Nonmetals
Fluorine
Chlorine
Bromine
Iodine

Polyatomic Ions

Name	Formula
Ammonium	NH ₄ ⁺
Acetate	C ₂ H ₃ O ₂ ⁻
Carbonate	CO ₃ ⁻²
Bicarbonate or Hydrogen carbonate	HCO ₃ ⁻
Hypochlorite	ClO ⁻
Chlorite	ClO ₂ ⁻
Chlorate	ClO ₃ ⁻
Perchlorate	ClO ₄ ⁻
Hydroxide	OH ⁻
Cyanide	CN ⁻
Nitrate	NO ₃ ⁻
Nitrite	NO ₂ ⁻
Chromate	CrO ₄ ⁻²
Oxalate	C ₂ O ₄ ⁻²
Sulfate	SO ₄ ⁻²
Sulfite	SO ₃ ⁻²
Hydrogen sulfate	HSO ₄ ⁻
Phosphate	PO ₄ ⁻³
Phosphite	PO ₃ ⁻³
Hydrogen phosphate	HPO ₄ ⁻²

Metric Conversions

Unit	Abbr.	Meaning
Mega-	M	10 ⁶
kilo-	k	10 ³
hecto-	h	10 ²
deca-	da	10 ¹
Liter	L	1
gram	g	1
meter	m	1
deci-	d	10 ⁻¹
centi-	c	10 ⁻²
milli-	m	10 ⁻³
micro-	μ	10 ⁻⁶
nano-	n	10 ⁻⁹

Calorie Conversions

Water = 1 cal/ g°C
1000 cal = 1 Cal

High School Chemistry Formulas Cheat Sheet

David Baud



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Table of Contents High School Chemistry Formulas Cheat Sheet

1. Understanding the eBook High School Chemistry Formulas Cheat Sheet
 - The Rise of Digital Reading High School Chemistry Formulas Cheat Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying High School Chemistry Formulas Cheat Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an High School Chemistry Formulas Cheat Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from High School Chemistry Formulas Cheat Sheet
 - Personalized Recommendations
 - High School Chemistry Formulas Cheat Sheet User Reviews and Ratings
 - High School Chemistry Formulas Cheat Sheet and Bestseller Lists

5. Accessing High School Chemistry Formulas Cheat Sheet Free and Paid eBooks
 - High School Chemistry Formulas Cheat Sheet Public Domain eBooks
 - High School Chemistry Formulas Cheat Sheet eBook Subscription Services
 - High School Chemistry Formulas Cheat Sheet Budget-Friendly Options
6. Navigating High School Chemistry Formulas Cheat Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - High School Chemistry Formulas Cheat Sheet Compatibility with Devices
 - High School Chemistry Formulas Cheat Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of High School Chemistry Formulas Cheat Sheet
 - Highlighting and Note-Taking High School Chemistry Formulas Cheat Sheet
 - Interactive Elements High School Chemistry Formulas Cheat Sheet
8. Staying Engaged with High School Chemistry Formulas Cheat Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers High School Chemistry Formulas Cheat Sheet
9. Balancing eBooks and Physical Books High School Chemistry Formulas Cheat Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection High School Chemistry Formulas Cheat Sheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine High School Chemistry Formulas Cheat Sheet
 - Setting Reading Goals High School Chemistry Formulas Cheat Sheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of High School Chemistry Formulas Cheat Sheet
 - Fact-Checking eBook Content of High School Chemistry Formulas Cheat Sheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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