

Posttest (Stoichiometry)

Solve the following problems analytically. Attached your computations and observe correct significant figures in your final answer.

- Carbon dioxide and water vapor are produced from the combustion of propane gas.

Equation:



- What volume of oxygen is needed for the complete combustion of 0.650 L of propane?

Note: 22.4 L is the volume of the any gasses at STP.

Solution: Mass-Mass factors

$$0.650 \text{ L-C}_3\text{H}_8 \times \frac{1 \text{ mole-C}_3\text{H}_8}{22.4 \text{ L-C}_3\text{H}_8} \times \frac{5 \text{ mole-O}_2}{1 \text{ mole-C}_3\text{H}_8} \times \frac{22.4 \text{ L O}_2}{1 \text{ mole-O}_2} = 3.25 \text{ L O}_2$$

- What volume of carbon dioxide can be produced from the reaction?

$$0.650 \text{ L-C}_3\text{H}_8 \times \frac{1 \text{ mole-C}_3\text{H}_8}{22.4 \text{ L-C}_3\text{H}_8} \times \frac{3 \text{ mole-CO}_2}{1 \text{ mole-C}_3\text{H}_8} \times \frac{22.4 \text{ L CO}_2}{1 \text{ mole-CO}_2} = 1.95 \text{ L CO}_2$$

- Given the following unbalanced equation: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

(a.) What is the molar mass of each species?

Chemical Species	Solution	Molar Mass
$\text{N}_2(\text{g})$	2 moles of N x 14.01 g/mol = 28.02 g/mole	28.02 g/mole
$3\text{H}_2(\text{g})$	3(2) moles of H x 1.01g/mole = 6.06g/mole	6.06 g/mole
$2\text{NH}_3(\text{g})$	2(1) moles of N x 14.01g/mole = 28.02 g/mole 2(3) moles of H x 1.01 g/mole = 6.06 g/mole 28.02 g/mole + 6.06 g/mole = 34.08 g/mole	34.08 g/mole

(b) Determine the mole ratio of the 3 species?

Chemical Species	Mole Ratio
N_2 and H_2	1 mol. N_2 /3 mol. H_2 And 3 mol. H_2 /1 mol. N_2
N_2 and NH_3	1 mol. N_2 /2 mol. NH_3 And

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