



A titrimetric analysis requires the careful addition of titrant.

Experiment 9

A Volumetric Analysis

- To prepare and standardize a sodium hydroxide solution
- To determine the molar concentration of a strong acid

OBJECTIVES

The following techniques are used in the Experimental Procedure:

TECHNIQUES



A chemical analysis that is performed primarily with the aid of volumetric glassware (e.g., pipets, burets, volumetric flasks) is called a **volumetric analysis**. For a volumetric analysis procedure, a known quantity or a carefully measured amount of one substance reacts with a to-be-determined amount of another substance with the reaction occurring in aqueous solution. The volumes of all solutions are carefully measured with volumetric glassware.

The known amount of the substance for an analysis is generally measured and available in two ways:

1. As a **primary standard**—An accurate mass (and thus, moles) of a solid substance is measured on a balance, dissolved in water, and then reacted with the substance being analyzed.
2. As a **standard solution**—A measured number of moles of substance is present in a measured volume of solution, generally expressed as the molar concentration (or molarity) of the substance. A measured volume of the standard solution then reacts with the substance being analyzed.

The reaction of the known substance with the substance to be analyzed, occurring in aqueous solution, is generally conducted by a titration procedure.

The titration procedure requires a buret to dispense a liquid, called the **titrant**, into a flask containing the **analyte** (Figure 9.1a, page 134). For the acid–base titration studied in Part B of this experiment, the titrant is a standard solution of sodium hydroxide and the analyte is an acid.

INTRODUCTION

Primary standard: a substance that has a known high degree of purity, a relatively large molar mass, is nonhygroscopic, and reacts in a predictable way

Standard solution: a solution having a very well known concentration of a solute

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Table of Contents Ja Beran Lab Manual Answers

1. Understanding the eBook Ja Beran Lab Manual Answers
 - The Rise of Digital Reading Ja Beran Lab Manual Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ja Beran Lab Manual Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ja Beran Lab Manual Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ja Beran Lab Manual Answers
 - Personalized Recommendations
 - Ja Beran Lab Manual Answers User Reviews and Ratings
 - Ja Beran Lab Manual Answers and Bestseller Lists
5. Accessing Ja Beran Lab Manual Answers Free and Paid eBooks
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 - Ja Beran Lab Manual Answers eBook Subscription Services
 - Ja Beran Lab Manual Answers Budget-Friendly Options

6. Navigating Ja Beran Lab Manual Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ja Beran Lab Manual Answers Compatibility with Devices
 - Ja Beran Lab Manual Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ja Beran Lab Manual Answers
 - Highlighting and Note-Taking Ja Beran Lab Manual Answers
 - Interactive Elements Ja Beran Lab Manual Answers
8. Staying Engaged with Ja Beran Lab Manual Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ja Beran Lab Manual Answers
9. Balancing eBooks and Physical Books Ja Beran Lab Manual Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ja Beran Lab Manual Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ja Beran Lab Manual Answers
 - Setting Reading Goals Ja Beran Lab Manual Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ja Beran Lab Manual Answers
 - Fact-Checking eBook Content of Ja Beran Lab Manual Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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