



GRAM FORMULA MASS

Name _____

Determine the gram formula mass (the mass of one mole) of each compound below.

$$1. \text{KMnO}_4 \quad 39 + 55 + 16(4) = \boxed{158}$$

$$2. \text{KCl} \quad 39 + 35 = \boxed{74}$$

$$3. \text{Na}_2\text{SO}_4 \quad 23(2) + 32 + 16(4) = \boxed{142}$$

$$4. \text{Ca(NO}_3)_2 \quad 40 + (14 + 16 \cdot 3)2 = \boxed{164}$$

$$5. \text{Al}_2(\text{SO}_4)_3 \quad 27 + (32 + 16 \cdot 4)3 = \boxed{315}$$

$$6. (\text{NH}_4)_3\text{PO}_4 \quad (14 + 4)(3) + 31 + 16(4) = \boxed{149}$$

$$7. \text{CuSO}_4 \cdot 5\text{H}_2\text{O} \quad 64 + 32 + 16(4) + 18(5) = \boxed{250}$$

$$8. \text{Mg}_3(\text{PO}_4)_2 \quad 24(3) + (31 + 16 \cdot 4)2 = \boxed{262}$$

$$9. \text{Zn}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 2\text{H}_2\text{O} \quad 65 + [2(24 + 3) + 16(2)2 + 18(2)] = \boxed{219}$$

Gram Formula Mass Answer Key

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Table of Contents Gram Formula Mass Answer Key

1. Understanding the eBook Gram Formula Mass Answer Key
 - The Rise of Digital Reading Gram Formula Mass Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Gram Formula Mass Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gram Formula Mass Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gram Formula Mass Answer Key
 - Personalized Recommendations
 - Gram Formula Mass Answer Key User Reviews and Ratings
 - Gram Formula Mass Answer Key and Bestseller Lists
5. Accessing Gram Formula Mass Answer Key Free and Paid eBooks

- Gram Formula Mass Answer Key Public Domain eBooks
 - Gram Formula Mass Answer Key eBook Subscription Services
 - Gram Formula Mass Answer Key Budget-Friendly Options
6. Navigating Gram Formula Mass Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Gram Formula Mass Answer Key Compatibility with Devices
 - Gram Formula Mass Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gram Formula Mass Answer Key
 - Highlighting and Note-Taking Gram Formula Mass Answer Key
 - Interactive Elements Gram Formula Mass Answer Key
 8. Staying Engaged with Gram Formula Mass Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gram Formula Mass Answer Key
 9. Balancing eBooks and Physical Books Gram Formula Mass Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gram Formula Mass Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gram Formula Mass Answer Key
 - Setting Reading Goals Gram Formula Mass Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gram Formula Mass Answer Key
 - Fact-Checking eBook Content of Gram Formula Mass Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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