- b) Calculate the density of NaCl, if edge length of NaCl unit cell is 564 pm. [Molar mass of NaCl = 58.5 g/mol]. (2)
 [March 2013]
- 11) Unit cells can be broadly classified into 2 categories primitive and centred unit cells.
 - a) What is a unit cell? (1)
 - b) Name the three types of centred unit cells. (1 %).
 - c) The unit cell dimension of a particular crystal system is a = b = c and $\alpha = \beta = \gamma = 90^{\circ}$. Identify the crystal system.
 - d) Give one example for the above crystal system. (%) [SAY 2013]
- 12) a) Every substance has some magnetic properties associated with it. How will you account for the following magnetic properties?
 - Paramagnetic property (1)
 - ii) Ferromagnetic property (1)
 - b) A compound is formed by two elements P and Q. Atoms of Q (as anions) make hcp lattice and those of the element P (as cations) occupy all the tetrahedral voids. What is the formula of the compound? (2) [March 2014]
- a) Crystalline solids are 'anisotropic'. What is anisotropy? (1)
 - b) Copper crystals have fcc unit cells.
 - i) Compute the number of atoms per unit cell of copper crystals. (1)
 - ii) Calculate the mass of a unit cell of copper crystals. (Atomic mass of Cu = 63.54u) (2) [SAY 2014]
- Unit cells can be divided into two categories primitive and centred unit cells.
 - a) Differentiate between unit cell and crystal lattice. (2)
 - b) Calculate the number of atoms per unit cell in the following:
 - Body centred cubic unit cell (bcc)
 - Face centred cubic unit cell (fcc) (2) [March 2015]
- 15) a) Which of the following is not a characteristic of a crystalline solid?
 - i) Definite heat of fusion ii) isotropic nature iii) A regular orderly arrangement of constituent particles
 - iv) A true solid (1)
 - Frenkel defect and Schottky defects are two stoichiometric defects found in crystalline solids.
 - i) What are stoichiometric defects?
 - Write any two differences between Schottky and Frenkel defects? (2) [SAY 2015]
- 16) a) Which of the following is a molecular solid?
 - (a) Diamond (b) graphite (c) ice (d) quartz (1)
 - (b) Unit cells can be classified into primitive and centered unit cells. Differentiate between primitive and centered unit cells. (1)
- (c) Presence of excess Na makes NaCl crystal coloured. Explain on the basis of crystal defects. (2) [March 2016]
- 17) A unit cell is a term related to crystal structure.
 - a) What do you mean by unit cell? (1)
 - b) Name any two types of cubic unit cells. (1)
 - c) Calculate the number of atoms in each of the above mentioned cubic unit cells. (1)
 - d) Identify the substance which shows Frenkel defect;
- (i) NaCl (ii) KCl (iii) ZnS (iv) AgBr (1) [SAY 2016]
- 18) a) Identify the non-stoichiometric defect
 - (i) Schottky defect (ii) Frenkel defect (iii) interstitial defect (iv) Metal deficiency defect (1)
 - What type of substance could make better permanent magnets ferrimagnetic or ferromagnetic? Justify your answer. (2)

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Wolfgang Guggemos

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- 1. Understanding the eBook Guide For Plus Two Chemistry
 - The Rise of Digital Reading Guide For Plus Two Chemistry
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Guide For Plus Two Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guide For Plus Two Chemistry
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Guide For Plus Two Chemistry
 - Personalized Recommendations

- Guide For Plus Two Chemistry User Reviews and Ratings
- Guide For Plus Two Chemistry and Bestseller Lists
- 5. Accessing Guide For Plus Two Chemistry Free and Paid eBooks
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 - Guide For Plus Two Chemistry Budget-Friendly Options
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 - Highlighting and Note-Taking Guide For Plus Two Chemistry
 - Interactive Elements Guide For Plus Two Chemistry
- 8. Staying Engaged with Guide For Plus Two Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guide For Plus Two Chemistry
- 9. Balancing eBooks and Physical Books Guide For Plus Two Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guide For Plus Two Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Guide For Plus Two Chemistry
 - Setting Reading Goals Guide For Plus Two Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Guide For Plus Two Chemistry
 - Fact-Checking eBook Content of Guide For Plus Two Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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