

Student:	Date:	
Topic:	Class:	
Main Ideas/Questions	Notes/Examples	
<b>Negative Exponents</b>	Negative exponents can be rewritten using positive exponents using the <b>NEGATIVE EXPONENT RULE</b> : <div> <math display="block">x^{-a} = \frac{1}{x^a}</math> </div>	
<b>Examples</b>	Directions: Rewrite each expression using positive exponents. <div>             1. <math>x^{-3}</math>             2. <math>-3a^{-2}</math>             3. <math>-7ab^{-3}</math> </div>	
	Directions: Simplify each expression. Make sure that answers contain only positive exponents. (Hint: use the rules, then move variables at the end!)	
	4. $x^3 \cdot x^{-5}$	5. $4x^2a^{-3} \cdot 5x^{-4}a^{-2}$
	6. $6a^{-2} \cdot -3a^{-7}$	7. $(a^2b^3c^{-4})(a^3b^{-2}c^5)$
	8. $(3p^2)^{-3}$	9. $(5q^2)^{-4}$
	10. $(2a^{-3})^{-4}$	11. $(8ab^2c)^{-3}$
	12. $\frac{b^2}{b^{-3}}$	13. $\frac{c^{-2}d^{-3}}{c^2d^{-4}}$
	14. $\frac{14a^2}{7a^{-3}}$	15. $\frac{-20a^2b}{2a^3b^{-2}}$

# Gina Wilson Homework 4 Negative Exponents

**L Towne**



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the quantity of heat for a process is represented by the letter q the sign of q for an endothermic process is positive because the system is gaining heat a chemical reaction or physical

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the quantities commonly encountered in chemical problems include the number of moles of a substance the number of atoms molecules or formula units of a substance and the mass in

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it is always a good idea to estimate the answer before you do the actual calculation in this case the mass given 35 00 g is less than the molar mass so the answer should be less than 1

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volume amount number of by counting a the

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the atomic mass of an element expressed in grams is the mass of a mole of the element the mass of a mole of an element is the molar mass to calculate the molar mass of a compound

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the amount of a substance that contains  $6.02 \times 10^{23}$  representative particles of that substance the amount of a substance that contains  $6.02 \times 10^{23}$  representative particles of that

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