

- Which event must *always* occur for a chemical reaction to take place?
 - formation of a precipitate
 - formation of a gas
 - effective collisions between reacting particles
 - addition of a catalyst to the reaction system
- In most aqueous reactions as temperature increases, the effectiveness of collisions between reacting particles
 - decreases
 - increases
 - remains the same
 - increases
- A 5.0-gram sample of zinc and a 50.-milliliter sample of hydrochloric acid are used in a chemical reaction. Which combination of these samples has the fastest reaction rate?
 - a zinc strip and 1.0 M HCl(aq)
 - a zinc strip and 3.0 M HCl(aq)
 - zinc powder and 1.0 M HCl(aq)
 - zinc powder and 3.0 M HCl(aq)
- A 1.0-gram piece of zinc reacts with 5 milliliters of HCl(aq). Which of these conditions of concentration and temperature would produce the greatest rate of reaction?
 - 1.0 M HCl(aq) at 20.°C
 - 1.0 M HCl(aq) at 40.°C
 - 2.0 M HCl(aq) at 20.°C
 - 2.0 M HCl(aq) at 40.°C
- At STP, which 4.0-gram zinc sample will react fastest with dilute hydrochloric acid?
 - lump
 - bar
 - powdered
 - sheet metal
- Given the balanced equation representing a reaction:

$$\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{H}_2\text{(g)} + \text{ZnCl}_2\text{(aq)}$$

Which set of reaction conditions produces $\text{H}_2\text{(g)}$ at the fastest rate?

- a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 20.°C
 - a 1.0-g lump of Zn(s) in 50. mL of 0.5 M HCl(aq) at 30.°C
 - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 20.°C
 - 1.0 g of powdered Zn(s) in 50. mL of 1.0 M HCl(aq) at 30.°C
- For a given reaction, adding a catalyst increases the rate of the reaction by
 - providing an alternate reaction pathway that has a higher activation energy
 - providing an alternate reaction pathway that has a lower activation energy
 - using the same reaction pathway and increasing the activation energy
 - using the same reaction pathway and decreasing the activation energy

- Given the balanced equation representing a reaction:



Which statement is true about energy in this reaction?

- The reaction is exothermic because it releases heat.
 - The reaction is exothermic because it absorbs heat.
 - The reaction is endothermic because it releases heat.
 - The reaction is endothermic because it absorbs heat.
- Which balanced equation represents an endothermic reaction?
 - $\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)}$
 - $\text{CH}_4\text{(g)} + 2\text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)} + 2\text{H}_2\text{O(g)}$
 - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
 - $\text{N}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO(g)}$
 - Given the balanced equation:



Which statement best describes this process?

- It is endothermic and entropy increases.
 - It is endothermic and entropy decreases.
 - It is exothermic and entropy increases.
 - It is exothermic and entropy decreases.
- Which change is exothermic?
 - freezing of water
 - melting of iron
 - vaporization of ethanol
 - sublimation of iodine
 - Which reaction releases the greatest amount of energy per 2 moles of product?
 - $2\text{CO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO}_2\text{(g)}$
 - $4\text{Al(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Al}_2\text{O}_3\text{(s)}$
 - $2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{H}_2\text{O(g)}$
 - $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)}$
 - Given the balanced equation representing a reaction at 101.3 kPa and 298 K:

$$\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{NH}_3\text{(g)} + 91.8 \text{ kJ}$$

Which statement is true about this reaction?

 - It is exothermic and ΔH equals -91.8 kJ.
 - It is exothermic and ΔH equals +91.8 kJ.
 - It is endothermic and ΔH equals -91.8 kJ.
 - It is endothermic and ΔH equals +91.8 kJ.
 - Which expression represents the ΔH for a chemical reaction in terms of the potential energy, PE , of its products and reactants?
 - PE of products + PE of reactants
 - PE of products - PE of reactants
 - PE of products \times PE of reactants
 - PE of products \div PE of reactants

Kinetics And Equilibrium Worksheet Answers

**John Ross, PURDUE UNIV LAFAYETTE
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Chemical Equilibrium and Solutions James E. Banks, 1967 Chemistry, Kinetics and Thermodynamics Stanford University. Department of Operations Research. Operations Research House, 1970 In the paper the equivalence of three ways of viewing the classical chemical equilibrium problem is discussed From the point of view of classical chemistry this problem can be stated as follows Given the set of reactions that may occur among all the species of a chemical system their equilibrium constants and the initial composition of the system i.e. the amounts of each species present in the system initially find the equilibrium composition of the system From the point of view of kinetics the problem is to determine the evolution of the various parameters describing the system for the evolution of the various parameters describing the system from their initial values When the parameters cease to change the system must be in equilibrium Given the same system thermodynamics seeks that composition at which the usable or free energy remaining in the system is at a minimum It can be shown that the system will then be at equilibrium Much of the material contained herein is old It is presented here in a consistent notation and in a way intended to emphasize the common aspects of the three points of view Author

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