

9. Each of four test tubes contains a different concentration of HCl(aq) at 25°C . A 1-gram cube of Zn is added to each test tube. In which test tube is the reaction occurring at the slowest rate?

A) 1.0 M HCl(aq)



B) 0.1 M HCl(aq)



C) 0.01 M HCl(aq)



D) 0.001 M HCl(aq)



10. As the concentration of reacting particles increases, the rate of reaction generally

- A) decreases
- B) increases
- C) remains the same

11. Which change would most likely decrease the rate of a chemical reaction?

- A) decreasing a reactant's concentration
- B) decreasing a reactant's surface area
- C) cooling the reaction mixture
- D) adding a catalyst to the reaction mixture

12. Which statement best describes how a catalyst increases the rate of a reaction?

- A) The catalyst provides an alternate reaction pathway with a higher activation energy.
- B) The catalyst provides an alternate reaction pathway with a lower activation energy.
- C) The catalyst provides the same reaction pathway with a higher activation energy.
- D) The catalyst provides the same reaction pathway with a lower activation energy.

13. If the pressure on gaseous reactants is increased, the rate of reaction is increased because there is an increase in the

- A) activation energy
- B) volume
- C) concentration
- D) heat of reaction

14. A 1.0-gram piece of Zn reacts with 1 milliliter of HCl(aq) . Which of these conditions of concentration and temperature would produce the greatest rate of reaction?

- A) 1.0 M HCl(aq) at 25°C
- B) 1.0 M HCl(aq) at 45°C
- C) 2.0 M HCl(aq) at 25°C
- D) 2.0 M HCl(aq) at 45°C

15. In a biochemical reaction, an enzyme acts as a catalyst. Which the

- A) activation energy of the reaction to decrease
- B) potential energy of the reactants to decrease
- C) kinetic energy of the reactants to increase
- D) heat of reaction to increase

16. Which reaction



As increase in the concentration of Fe(s) will

- A) decrease the production of $\text{FeCl}_2\text{(aq)}$
- B) decrease the frequency of collisions between Fe(s) and $\text{CuCl}_2\text{(aq)}$
- C) increase the production of $\text{FeCl}_2\text{(aq)}$
- D) increase the frequency of collisions between Fe(s) and Fe(s)

17. A reaction is most likely to occur when reactant particles collide with

- A) proper energy, only
- B) proper orientation, only
- C) both proper energy and proper orientation
- D) neither proper energy nor proper orientation

18. Given the reaction



The reaction will most likely occur at the greatest rate if A and B represent

- A) ionic molecular compounds in the solid phase
- B) ionic compounds in the solid phase
- C) volatile or nonpolar molecular compounds
- D) volatile or ionic compounds

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RC Schank



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