

PHYSICAL: Kinetic Molecular Theory

Learning Objectives

- At standard atmospheric conditions, all gases have approximately the same kinetic energy, and are made up of molecules that behave as independent particles (the postulates of the kinetic molecular theory).
- At standard atmospheric conditions, all gases have the same average kinetic energy.
- At standard atmospheric conditions, all gases have the same average molecular mass.

Model 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in three different phases of matter.



Figure 1: Representation of particles in different phases of matter.

Model 1 represents the three phases of matter. The particles in the gas phase are widely spaced and move rapidly. The particles in the liquid phase are closely packed and move more slowly. The particles in the crystalline solid phase are closely packed and move very slowly.

Key Concepts

1. Model 1 shows the three phases of matter. The particles in the gas phase are widely spaced and move rapidly. The particles in the liquid phase are closely packed and move more slowly. The particles in the crystalline solid phase are closely packed and move very slowly.

Table 1: Representation of particles in different phases of matter.

	Gas	Liquid	Crystalline Solid
Particle arrangement	Particles are widely spaced and move rapidly. The particles are randomly distributed.	Particles are closely packed and move more slowly. The particles are randomly distributed.	Particles are closely packed and move very slowly. The particles are in a highly ordered, repeating lattice structure.
Particle movement	Particles move rapidly in all directions. The particles are in constant motion.	Particles move more slowly in all directions. The particles are in constant motion.	Particles move very slowly in all directions. The particles are in constant motion.

Kinetic Molecular Theory Pogil Answer

J Elliott



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