

# 7&8 • Atomic Structure & Periodicity

## PRACTICE TEST

$$Rhc = A$$

$$A = 2.18 \times 10^{-18} \text{ J}$$

$$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$

$$R = 1.097 \times 10^7 \text{ m}^{-1}$$

$$c = 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}$$

$$\text{mass of an electron} = 9.11 \times 10^{-31} \text{ kg}$$

1. What wavelength corresponds to a frequency of  $8.22 \times 10^9 \text{ Hz}$ ?

$$\lambda = \frac{c}{\nu}$$

- a) 0.307 m                      d) 0.110 m

- b) 0.0365 m                      c) 27.4 m

- c) 0.122 m                       $\frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{8.22 \times 10^9 \text{ s}^{-1}} = 0.036496 \text{ m}$

2. A radio station transmits at 110 MHz ( $110 \times 10^6 \text{ Hz}$ ). What wavelength is this radio wave?

- a)  $3.65 \times 10^{-5} \text{ m}$                       c)  $3.81 \times 10^{-5} \text{ m}$

- b) 3.30 m                      d) 2.73 m

$$\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{110 \times 10^6 \text{ s}^{-1}} = 2.73 \text{ m}$$

3. Which one of the following is NOT a proper unit for frequency?

- a) Hz

- b)  $\text{s}^{-1}$

- c)  $\text{m}\cdot\text{s}^{-1}$  velocity

- d)  $\frac{1}{\text{sec}}$

4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.

$$n=6 \rightarrow n=2$$

$$E_6 = -2.18 \times 10^{-18} \text{ J} - E_2 = -2.18 \times 10^{-18} \text{ J}$$

- a) 0.12334 m

- d)  $4.1029 \times 10^{-7} \text{ m}$

- b) 24.373 m

- e) 36.559 m

- c)  $2.7353 \times 10^{-7} \text{ m}$

$$E = \frac{hc}{\lambda} \quad \lambda = \frac{hc}{E} = 4.1 \times 10^{-7} \text{ m}$$

$$\Delta E = 4.84 \times 10^{-19} \text{ J}$$

5. What is the relationship between the energy of a photon of light and its frequency?

- a)  $E = \nu$

- d)  $E = \frac{1}{h\nu}$

- b)  $E = \frac{h}{\nu}$

- e)  $E = \frac{\nu}{h}$

- c)  $E = h\nu$

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?

$$\Delta E = E_3 - E_2$$

- a)  $1.52 \times 10^{-4} \text{ J}$

- d)  $4.48 \times 10^{-19} \text{ J}$

- b)  $3.63 \times 10^{-19} \text{ J}$

- e)  $3.03 \times 10^{-19} \text{ J}$

- c)  $2.18 \times 10^{-19} \text{ J}$

$$-\frac{2.18 \times 10^{-18} \text{ J}}{9} - \left( -\frac{2.18 \times 10^{-18} \text{ J}}{4} \right)$$

7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.

- a)  $3.03 \times 10^{-12} \text{ m}$

- c)  $3.30 \times 10^{11} \text{ m}$

- b)  $2.42 \times 10^{-12} \text{ m}$

- d)  $1.59 \times 10^{-25} \text{ m}$

$$\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{9.11 \times 10^{-31} \text{ kg} \times 0.8 \times 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}$$

8. What resultant is expected from the interference of the two waves shown below?



- a)

- c)

- b)

- d)

9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?

- a)  $n$

- c)  $m_l$

- b)  $\ell$

- d)  $m_s$

# Gpb Chemistry Answers Atomic Structure

**S Marginson**



## **Gpb Chemistry Answers Atomic Structure:**

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We will explore periodic trends such as atomic radius ionization energy and electron affinity and discuss how these trends are related to an element's position on the table. In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding stoichiometry and the behavior of gases. We will also explore the importance of understanding atomic structure in fields such as biochemistry environmental science and materials science. Throughout this book we will provide clear explanations helpful illustrations and interactive activities to reinforce key concepts. Whether you are a visual learner a hands on learner or a theoretical learner this book is designed to meet your needs. We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry. So let's dive in and unlock the mysteries of the atom together.

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