

ex: The pH of blood is 7.41. Find the pOH, $[H^+]$, and $[OH^-]$.

$$pH + pOH = 14.00$$

$$7.41 + pOH = 14.00$$

$$pOH = 6.59$$



$$pOH = -\log [OH^-]$$

$$-pOH = \log [OH^-]$$

$$10^{-pOH} = 10^{\log [OH^-]}$$

$$10^{-pOH} = [OH^-]$$

$$pH = -\log [H^+]$$

$$-pH = \log [H^+]$$

$$10^{-pH} = 10^{\log [H^+]}$$

$$10^{-pH} = [H^+]$$

"10 to the log of something equals that something."

$$10^{-7.41} = [H^+] = 3.9 \times 10^{-8} M$$

→ punch in to calculator to get ↗

$$10^{-6.59} = [OH^-] = 2.6 \times 10^{-7} M$$

* Calculating the pH of Strong Acid Solutions.

↳ HCl , HNO_3 , H_2SO_4 , $HClO_4$

- it is essential to know the major components (species) in the solution, and to know the minor species that can be ignored.

ex: $[HNO_3]$ → a very strong acid, and strong acids dissociate completely in solution.

↳ so, a 1.0M HNO_3 solution contains no HNO_3 !! Instead it contains H^+ , NO_3^- , and H_2O .

ex: Calculate the pH of a 0.10M HCl solution.



so 0.10M H^+

↳ H^+ , Cl^- , H_2O (no " HCl ")

↳ its " H^+ " contribution is negligible.
($1.0 \times 10^{-7} M H^+$)

$$pH = -\log [H^+] = -\log (0.10) \Rightarrow pH = 1.00$$

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Gabriella Székely

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Gpb Chemistry 1502 Notes Answer

1. Understanding the eBook Gpb Chemistry 1502 Notes Answer
 - The Rise of Digital Reading Gpb Chemistry 1502 Notes Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Gpb Chemistry 1502 Notes Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gpb Chemistry 1502 Notes Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gpb Chemistry 1502 Notes Answer
 - Personalized Recommendations
 - Gpb Chemistry 1502 Notes Answer User Reviews and Ratings
 - Gpb Chemistry 1502 Notes Answer and Bestseller Lists

5. Accessing Gpb Chemistry 1502 Notes Answer Free and Paid eBooks
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 - Gpb Chemistry 1502 Notes Answer eBook Subscription Services
 - Gpb Chemistry 1502 Notes Answer Budget-Friendly Options
6. Navigating Gpb Chemistry 1502 Notes Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Gpb Chemistry 1502 Notes Answer Compatibility with Devices
 - Gpb Chemistry 1502 Notes Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gpb Chemistry 1502 Notes Answer
 - Highlighting and Note-Taking Gpb Chemistry 1502 Notes Answer
 - Interactive Elements Gpb Chemistry 1502 Notes Answer
8. Staying Engaged with Gpb Chemistry 1502 Notes Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gpb Chemistry 1502 Notes Answer
9. Balancing eBooks and Physical Books Gpb Chemistry 1502 Notes Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gpb Chemistry 1502 Notes Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gpb Chemistry 1502 Notes Answer
 - Setting Reading Goals Gpb Chemistry 1502 Notes Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gpb Chemistry 1502 Notes Answer
 - Fact-Checking eBook Content of Gpb Chemistry 1502 Notes Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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