

Activity A:

Subatomic particles

Get the Gizmo ready:

- Use the arrows to create an atom with two protons, two neutrons, and two electrons.
- Turn on **Show element name**.



Question: What are the properties of protons, neutrons, and electrons?

1. **Observe:** Turn on **Show element symbol** and **Element notation**. Three numbers surround the element symbol: the **mass number** (A), electrical charge (no number is displayed if the atom is neutral), and the **atomic number** (Z).
2. **Investigate:** Watch how the numbers change as you add or remove particles.



- A. Which number is equal to the number of protons in the atom?
- B. How can you calculate the number of neutrons (N) in an atom?
- C. Which particle (proton, neutron, or electron) has a
Positive charge?
Negative charge?

neutrons

by the number of protons

electron

proton

neutron

3. **Observe:** Turn on **Isotope**. You can observe how the number of protons and neutrons in an atom can vary. The number of protons is the same for all atoms of an element, but the number of neutrons can vary. This creates different isotopes of the same element. For example, Helium has two isotopes: Helium-3 and Helium-4. Helium-3 has 2 protons and 1 neutron, while Helium-4 has 2 protons and 2 neutrons.

- A. What are the mass numbers of Helium-3 and Helium-4?
- B. What are the mass numbers of Lithium-6 and Lithium-7?
- C. Carbon has two isotopes: Carbon-12 and Carbon-13. What are their mass numbers?

3 and 4

6 and 7

12 and 13

4. **Observe:** Turn on **Electron** to observe the following questions.

- A. How many electrons are in a neutral atom of Helium?
- B. How many electrons are in an atom of Hydrogen?
- C. What is the mass number of an atom with 2 protons and 1 neutron?
- D. What is the mass number of an atom with 2 protons and 2 neutrons?
- E. How many electrons are in Helium?

2

1

3

4

2

Gizmo Worksheet Answers Element Builder

Huangqi Zhang

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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Table of Contents Gizmo Worksheet Answers Element Builder

1. Understanding the eBook Gizmo Worksheet Answers Element Builder
 - The Rise of Digital Reading Gizmo Worksheet Answers Element Builder
 - Advantages of eBooks Over Traditional Books
2. Identifying Gizmo Worksheet Answers Element Builder
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gizmo Worksheet Answers Element Builder
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gizmo Worksheet Answers Element Builder
 - Personalized Recommendations
 - Gizmo Worksheet Answers Element Builder User Reviews and Ratings
 - Gizmo Worksheet Answers Element Builder and Bestseller Lists
5. Accessing Gizmo Worksheet Answers Element Builder Free and Paid eBooks
 - Gizmo Worksheet Answers Element Builder Public Domain eBooks
 - Gizmo Worksheet Answers Element Builder eBook Subscription Services
 - Gizmo Worksheet Answers Element Builder Budget-Friendly Options
6. Navigating Gizmo Worksheet Answers Element Builder eBook Formats

- ePub, PDF, MOBI, and More
- Gizmo Worksheet Answers Element Builder Compatibility with Devices
- Gizmo Worksheet Answers Element Builder Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gizmo Worksheet Answers Element Builder
 - Highlighting and Note-Taking Gizmo Worksheet Answers Element Builder
 - Interactive Elements Gizmo Worksheet Answers Element Builder
- 8. Staying Engaged with Gizmo Worksheet Answers Element Builder
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gizmo Worksheet Answers Element Builder
- 9. Balancing eBooks and Physical Books Gizmo Worksheet Answers Element Builder
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gizmo Worksheet Answers Element Builder
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gizmo Worksheet Answers Element Builder
 - Setting Reading Goals Gizmo Worksheet Answers Element Builder
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gizmo Worksheet Answers Element Builder
 - Fact-Checking eBook Content of Gizmo Worksheet Answers Element Builder
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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