

4. **Predict:** What do you think will happen if you fire a 7-eV photon at the atom of **Gas A**? How about a 13-eV or a 19-eV photon?

The electron will move to orbitals that are farther away.

5. **Gather data:** Test your predictions with the Gizmo and fill in the table below. (The first row has been filled in for you.)

Photon energy	Effect on electron	Energy of emitted photon(s)
4 eV	Electron moves up to $n = 2$ and then back down to $n = 1$.	4 eV
7 eV	The electron climbs to $n=3$ before descending to $n=1$.	4 eV and 3 eV
13 eV	The electron climbs to $n=4$ before descending to $n=1$.	4 eV and 9 eV
19 eV	The electron climbs to $n=5$ before descending to $n=1$.	4 eV and 15 eV

6. **Analyze:** Find the total energy of each set of emitted photons. How does each sum relate to the energy of the absorbed photon?

Yes, it has only been divided.

7. **Explore:** With the **Energy (eV)** set to 19 eV, click **Fire** six times. Record the energy of the emitted photons each time. Record the results of each trial below.

Trial	Energy of emitted photons	Trial	Energy of emitted photons
1	3 eV, 4 eV, 6 eV and 6 eV	4	3 eV, 4 eV, 6 eV and 6 eV
2	6 eV and 13 eV	5	3 eV, 4eV, 12 eV
3	3 eV, 4 eV, 6 eV and 6 eV	6	4 eV and 15 eV

8. **Analyze:** When an electron moves from a higher orbit to a lower one, does it always follow the same path? Explain.

No, it has multiple ways to distribute the energy used to move it to the higher orbital.

Gizmo Bohr Model Of Hydrogen Answers Key

R Barnett

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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