

CHEMISTRY

CHEAT SHEET

CHEMICAL KINETICS

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DEFINITIONS:

- Rate of Reaction:** The change in concentration of a reactant or product per unit time.
- Rate Law:** An equation that relates the rate of a reaction to the concentration of the reactants.
- Order of Reaction:** The sum of the powers of the concentration terms in the rate law.
- Half-life:** The time required for the concentration of a reactant to decrease to half its initial value.

FACTORS AFFECTING RATE:

- Temperature
- Concentration
- Surface Area
- Catalysts

Rate of Reaction:

For a reaction: $aA + bB \rightarrow cC + dD$

Rate = $-\frac{1}{a} \frac{d[A]}{dt} = -\frac{1}{b} \frac{d[B]}{dt} = \frac{1}{c} \frac{d[C]}{dt} = \frac{1}{d} \frac{d[D]}{dt}$

Rate Law:

For a reaction: $aA + bB \rightarrow cC + dD$

Rate = $k[A]^m[B]^n$

Where k is the rate constant, $[A]$ and $[B]$ are the concentrations of the reactants, and m and n are the orders of reaction.

Order of Reaction:

For a reaction: $aA + bB \rightarrow cC + dD$

Order = $m + n$

Half-life:

For a reaction: $aA + bB \rightarrow cC + dD$

Half-life = $\frac{\ln 2}{k}$

Reaction Orders

Order	Rate Law	Half-life	Graph
Zero-order	Rate = $k[A]$	$t_{1/2} = \frac{[A]_0}{k}$	Graph of $[A]$ vs t is a straight line with a negative slope.
First-order	Rate = $k[A]$	$t_{1/2} = \frac{\ln 2}{k}$	Graph of $\ln[A]$ vs t is a straight line with a negative slope.
Second-order	Rate = $k[A]^2$	$t_{1/2} = \frac{1}{k[A]_0}$	Graph of $\frac{1}{[A]}$ vs t is a straight line with a positive slope.

Graphs of Chemical Kinetics

Graph of $[A]$ vs t :

The concentration of a reactant decreases exponentially over time.

Graph of $\ln[A]$ vs t :

The natural logarithm of the concentration of a reactant decreases linearly over time.

Graph of $\frac{1}{[A]}$ vs t :

The reciprocal of the concentration of a reactant increases linearly over time.

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PT Brinkman

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the PT Brinkman text.

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